

Chemistry Notes for class 12 Chapter 5

Surface Chemistry

Surface Chemistry is the branch of chemistry which deals with the phenomenon that occurs on the surfaces or interfaces, such phenomenon includes corrosion, catalysis, crystallisation, etc

Adsorption

Due to unbalanced attraction forces, accumulation of molecular species at the surface rather than in the bulk of a solid or liquid is termed as adsorption. The molecular species accumulates at the surface is termed as adsorbate and the material on the surface of which the adsorption takes place is called adsorbent, e.g.,

- (i) O_2 , H_2 , Cl_2 , NB_3 gases are adsorbed on the surface of charcoal.
- (ii) Silica gels adsorb water molecules from air.

Charcoal, silica gel, metals such as Ni, Cu, Ag, Pt and colloids are some adsorbents.

Important Characteristics of Adsorption

- 1. It is specific and selective in nature.
- 2. Adsorption is spontaneous process, therefore change in free energy (ΔG) is negative.

$$\Delta G = \Delta H - T\Delta S,$$

For the negative value of ΔG , in a system, in which randomness decreases, ΔH must be negative. Hence, adsorption is always exothermic.

Adsorption of hydrogen over Pt is called **occlusion**.

Desorption

It is a process of removing an adsorbed substance from a surface on which it is adsorbed, is known as desorption.

Distinction between Adsorption and Absorption

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 - The Rise of Digital Reading Download Surface Chemistry Class 12 Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Download Surface Chemistry Class 12 Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Download Surface Chemistry Class 12 Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Download Surface Chemistry Class 12 Notes
 - Personalized Recommendations
 - Download Surface Chemistry Class 12 Notes User Reviews and Ratings
 - Download Surface Chemistry Class 12 Notes and Bestseller Lists
5. Accessing Download Surface Chemistry Class 12 Notes Free and Paid eBooks

- Download Surface Chemistry Class 12 Notes Public Domain eBooks
 - Download Surface Chemistry Class 12 Notes eBook Subscription Services
 - Download Surface Chemistry Class 12 Notes Budget-Friendly Options
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 - Download Surface Chemistry Class 12 Notes Enhanced eBook Features
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- Adjustable Fonts and Text Sizes of Download Surface Chemistry Class 12 Notes
 - Highlighting and Note-Taking Download Surface Chemistry Class 12 Notes
 - Interactive Elements Download Surface Chemistry Class 12 Notes
8. Staying Engaged with Download Surface Chemistry Class 12 Notes
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Download Surface Chemistry Class 12 Notes
9. Balancing eBooks and Physical Books Download Surface Chemistry Class 12 Notes
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Download Surface Chemistry Class 12 Notes
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Download Surface Chemistry Class 12 Notes
- Setting Reading Goals Download Surface Chemistry Class 12 Notes
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Download Surface Chemistry Class 12 Notes
- Fact-Checking eBook Content of Download Surface Chemistry Class 12 Notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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