

Gaseous State

1. GASEOUS STATE

When the molecular forces of attraction between the particles of a matter are minimum, the particles exist in a state known as gaseous state or a state of matter in which molecules are far away from each other and free to move in available space is called gaseous state.

1.1 Properties of gases

- (a) They do not have definite shape and volume.
- (b) They can occupy whole space open to them.
- (c) Gases have unlimited dispensability and high compressibility.
- (d) They have very low densities because of negligible intermolecular forces.
- (e) Gases exert pressure on the wall of the container with perfectly elastic collisions.
- (f) They diffuse rapidly through each other to form homogenous mixture against the electric, magnetic and gravitational field.

1.2 Parameters of Gases

The characteristics of gases are described in terms of four measurable parameters and it is also called as measurable properties of gases which are

- (a) Mass (b) Volume (c) Pressure and (d) Temperature

(a) **Mass (m)** – The mass of a gas is denoted by 'm' which is related to the no of moles 'n'.

$$\text{Therefore, } n \text{ (no of moles)} = \frac{m(\text{mass in grams})}{M(\text{Molar mass})}$$

$$\text{so, } m = n \times M$$

(b) **Volume V** -

- (i) Gases occupy whole space available to them. The volume occupied by a gas is simply the volume of container in which it is filled.
- (ii) The volume of a gas is denoted by 'V' and it is measured in units of litre or cubic metre (m^3) or cm^3 or dm^3 .
- (iii) $1 \text{ litre} = 1 \text{ dm}^3 = 1000 \text{ cm}^3 = 1000 \text{ ml}$

(c) **Pressure (P)** -

- (i) It is force acting per unit area. A confined gas exerts uniform pressure on the walls of its container in all the direction.
- (ii) It is denoted by 'P' and specified in pascal (P_a)
- (iii) Other units of pressure are atm, cm, hg, mmHg, N/m^2 , bar, torr.
- (iv) $1 \text{ atm} = 76 \text{ cm Hg} = 760 \text{ mm Hg} = 1.013 \times 10^5 \text{ N/m}^2 = 1.013 \times 10^5 P_a = 1.013 \text{ bar} = 760 \text{ torr}$.

$$(v) P \text{ (Pressure)} = \frac{F(\text{Force})}{A(\text{Area})} = \frac{\text{Mass} \times \text{Acceleration}}{\text{Area}}$$

- (vi) Pressure exerted by a gas is due to kinetic energy ($\text{K.E.} = \frac{1}{2}mv^2$) of the gases molecules.
- (vii) K.E. of the gas molecule increases, as the temperature is increased so, pressure of a gas is directly proportional to temperature. $P \propto T$.

(d) **Temperature (T)** -

- (i) The temperature of a gas is denoted by 'T' and it is measured in the unit of kelvin (K).
- (ii) Other units of temperature are $^{\circ}\text{C}$, $^{\circ}\text{F}$, $^{\circ}\text{R}$
- (iii) $K = ^{\circ}\text{C} + 273.15$

$$(iv) \frac{x^{\circ}\text{C}}{5} = \frac{(y^{\circ}\text{F} - 32)}{9}$$

2. GAS LAWS

The certain laws which relate the four parameters are called gas laws.

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