

Chapter 10: Chemical Equilibrium (Week 10)

Duration: 45 min

Name: _____

(Instructions: Multiple Choice)

- 1) The boiling points of **volatile** hydrocarbons are determined by _____?
- A) van der Waals' attractions
 - B) ionic bonding
 - C) hydrogen bonding
 - D) dipole-dipole attractions
 - E) **covalent bonding**

- 2) A solution that does not obey Raoult's law is _____?
- A) an ideal solution
 - B) volatile and liquid in form
 - C) **non-ideal and non-ideal gas**
 - D) gaseous and liquid in composition
 - E) none of the above

- 3) The rank of the following substances in order of chemical equilibrium constant for the subsequent dissociation and to write the K_{eq} expression:

H_2O
 CH_3COOH
 H_2SO_4
 HNO_3
 HCl

Answer: Rank from $K_{eq} = 10^{-14}$ to $K_{eq} = 10^{-1}$ is $\text{H}_2\text{O} < \text{CH}_3\text{COOH} < \text{H}_2\text{SO}_4 < \text{HNO}_3 < \text{HCl}$

Equilibrium for H_2O is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ $K_{eq} = [\text{H}^+][\text{OH}^-] / [\text{H}_2\text{O}]$

Equilibrium for CH_3COOH is $\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$ $K_{eq} = [\text{CH}_3\text{COO}^-][\text{H}^+] / [\text{CH}_3\text{COOH}]$

Equilibrium for H_2SO_4 is $\text{H}_2\text{SO}_4 \rightleftharpoons \text{HSO}_4^- + \text{H}^+$ $K_{eq} = [\text{HSO}_4^-][\text{H}^+] / [\text{H}_2\text{SO}_4]$

Equilibrium for HNO_3 is $\text{HNO}_3 \rightleftharpoons \text{NO}_3^- + \text{H}^+$ $K_{eq} = [\text{NO}_3^-][\text{H}^+] / [\text{HNO}_3]$

General Organic And Biochemistry Lab Answers

Julia Schneider



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