

Carbonate Alkalinity

- Consider a solution containing a fixed total dissolved concentration of CaCO_3 . At the CO_2 equivalence point, H_2CO_3 is the principal carbonate species, so we choose our components as H^+ , H_2O , CO_2 , and Ca^{2+} .



- The proton mole balance equation is then:

$$\text{TOT} = [\text{H}^+] - [\text{HCO}_3^-] - 2[\text{CO}_3^{2-}] - [\text{OH}^-]$$

- The alkalinity is then:

$$\text{Alk} = -\text{TOT} = -[\text{H}^+] + [\text{HCO}_3^-] + 2[\text{CO}_3^{2-}] + [\text{OH}^-]$$

- The above is also the *carbonate alkalinity: the alkalinity due to the presence of carbonate ions.*
 - In many solutions, carbonate alkalinity is very nearly equal total alkalinity. To be clear, we sometimes call the alkalinity as defined previously Total Alkalinity.

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Environmental Sampling and Analysis Maria Csuros, 2018-05-11
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