



- a. combination
b. decomposition

- c. displacement
d. exchange

19. Classify the following reaction.



- a. combination
b. decomposition

- c. displacement
d. exchange

20. The Roman numerals in the reaction given represent the coefficients in the balanced chemical equation. What are the values of the coefficients?



- I II III IV
a. 2 2 2 3
b. 1 1 1 2
c. 2 1 2 1
d. 1 2 2 1
e. 2 2 2 1

21. Which of the following is a diprotic acid?

- a. H_3PO_4
b. H_2O
c. Ca(OH)_2

- d. H_2SO_4
e. NH_3

22. Which compound can be classified as insoluble:

- a. $\text{Mg(ClO}_4\text{)}_2$;
b. $\text{Cu(NO}_3\text{)}_2$;
c. $(\text{NH}_4)_2\text{SO}_4$;

- d. Fe(OH)_3
e. NiSO_4 ;

23. Consider the reaction between 15.0 mL of a 1.00 M aqueous solution of AgNO_3 and 10.0 mL of a 1.00 M aqueous solution of K_2CO_3 . When these react, a precipitate is observed. What is present in solution after the reaction is complete (*Note: the precipitate is not considered to be in solution*)?

- a. Ag^+ , NO_3^- , K^+ , CO_3^{2-} , water
b. Ag^+ , NO_3^- , K^+ , water
c. K^+ , CO_3^{2-} , water

- d. NO_3^- , K^+ , CO_3^{2-} , water
e. water

24. Which statement about strong acids is true?

- a. Strong acids are weak electrolytes.
b. Strong acids are very concentrated.
c. Strong acids are almost entirely converted to ions when dissolved in water.
d. Acetic acid is a strong acid.
e. All of the above are true.

25. Which of the following is a weak electrolyte?

- a. HBr
b. Mg(OH)_2
c. NaCl

- d. HCN
e. KI

26. What is the correct formula for the ammonium ion?

- a. H_2O ;
b. H_3O^+ ;
c. OH^- ;

- d. H^+ ;
e. NH_4^+ ;

27. What is the oxidation state (number) of Cl in Cl_2O_5 ?

- a. +2
b. -2

- c. -3
d. -5

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