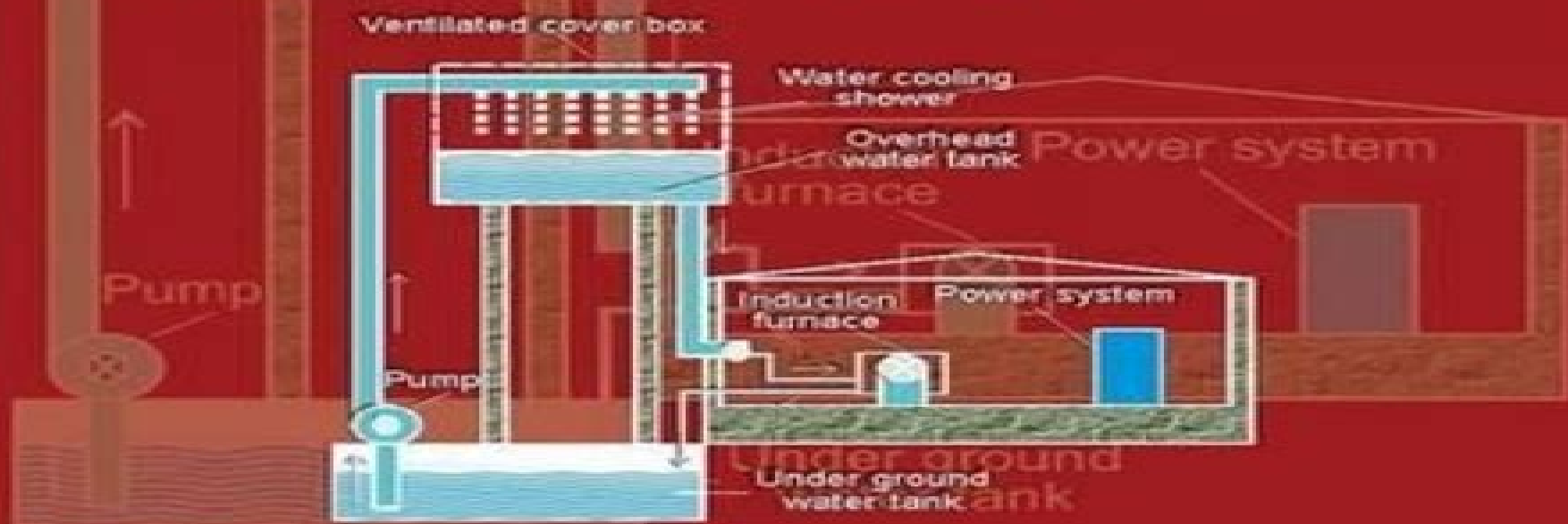


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Refractories and Pyrometry A V K Suryanarayana, 2015-10 Present day technology is vibrant and changing rapidly But the essential characteristics remain the same when a fuel is burnt the aim will always be to completely burn it and derive maximum heat out of it A furnace and its refractory linings are must to utilize the fuel When the fuel is burnt and some process s are performed in the furnace it becomes a consequential necessity to measure the temperature in the furnace to have a proper control over the operations An effort is made to give the students a deep insight into the utilization of fuels with some fundamentals essential to have a grasp of the subject This book thus tries to encompass the fuel utilization to a satisfactory level Salient features Units are converted to S I Units from CGS or FPS systems More material is added in Nuclear and Solar Energy topics **Metals Abstracts** ,1983 *Transactions* Iron and Steel Institute,1960 Ceramic Abstracts American Ceramic Society,1981 **Journal** ,1960 Includes the institute s Proceedings **Steelworks Fuels, Furnaces, Refractories and Instruments** George Reginald Bashforth,1960 *Fuels and Furnaces* ,1929 *Worked Examples in Heat Transfer, Fuels & Refractories, Fluid Flow in Furnace Technology* D. N. Gwyther,1985 **Fuel and Refractory Materials** Alexander Humboldt Sexton,1897 **Refractories and Furnaces** Francis Thompson Havard,1912 *Glass Furnace Refractories* Kaiser Refractories,1960 **Boiler Furnaces** UNKNOWN. AUTHOR,2015-06-12 Excerpt from Boiler Furnaces The process of combustion is affected by so many different physical conditions that in order to either design a furnace or superintend its operation it is necessary to fully understand the chemical changes which take place Among the physical conditions referred to above may be included strength of draft depth of fuel bod form of furnace and the various methods of firing employed So far as the generation of steam is concerned combustion may be considered as the chemical union of oxygen with the various elements of the fuel for which it has an affinity This union produces a definite amount of heat per pound of combustible depending upon the clement with which the oxygen combines The elements contained in the usual forms of fuel which enter into the process of combustion are oxygen carbon hydrogen and sulphur There are various other constituents present which have no fuel value such as the Iron silicon etc found in coal These usually exist in small quantities and are classed as impurities They produce a certain waste in the form of ash and in addition to this their temperature must be raised to that of the fire before becoming separated from the other elements and more or less of this heat Is lost as they are discharged from the fire Oxygen is the universal element of combustion it is an invisible gas and makes up about one fifth the volume of the air in an uncombined state It is also found in water being combined with hydrogen in this case and in coal and other fuels of vegetable origin combined with carbon and hydrogen It is usually present in coal in amounts varying from 1 to 25 per cent according to the grade Carbon is a solid and is found in a pure state In the form of graphite and charcoal In its crystalized state it forms diamonds It is also found In oils of various kinds and In tar combined with hydrogen Vegetable products of all kinds contain carbon In combination with oxygen and hydrogen and it is the principal heat producing element In coal and other fuels Including liquids and gases Hydrogen is a combustible gas and

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