

Chapter 10: Chemical Equilibrium (Week 10)

Duration: 45 min

Name: _____

(Instructions: Multiple Choice)

1) The boiling points of **volatile** hydrocarbons are determined by _____?

- A) van der Waals attractions
- B) ionic bonding
- C) hydrogen bonding
- D) dipole-dipole attractions
- E) **covalent bonding**

2) Water has a high heat capacity because it is _____?

- A) less dense than water
- B) volatile and hard to boil
- C) **hydrogen bond network**
- D) polar and water is abundant
- E) less dense than water

3) The acid of the following substances is used in chemical synthesis. Which has the highest dissociation constant and therefore the K_a expression:

- Acetic
- Chloroacetic
- Benzoic
- Phosphoric
- Hydrofluoric

Acetic acid $\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$ $K_a = 1.8 \times 10^{-5}$

Chloroacetic acid $\text{CH}_2\text{ClCOOH} \rightleftharpoons \text{CH}_2\text{ClCOO}^- + \text{H}^+$ $K_a = 1.4 \times 10^{-3}$

Benzoic acid $\text{C}_6\text{H}_5\text{COOH} \rightleftharpoons \text{C}_6\text{H}_5\text{COO}^- + \text{H}^+$ $K_a = 6.3 \times 10^{-5}$

Phosphoric acid $\text{H}_3\text{PO}_4 \rightleftharpoons \text{H}_2\text{PO}_4^- + \text{H}^+$ $K_a = 7.5 \times 10^{-3}$

Hydrofluoric acid $\text{HF} \rightleftharpoons \text{F}^- + \text{H}^+$ $K_a = 6.6 \times 10^{-4}$

General Organic And Biochemistry Lab Answers

William G. O'Neal



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National Library of Medicine Catalog National Library of Medicine (U.S.), 1966

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