

Physics

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(Chapter 13) (Kinetic Theory) (Additional Exercises)
(Class 11)

Question 13.14:

Given below are densities of some solids and liquids. Give rough estimates of the size of their atoms:

Substance	Atomic Mass (u)	Density (10^3 Kg m^{-3})
Carbon (diamond)	12.01	2.22
Gold	197.00	19.32
Nitrogen (liquid)	14.01	1.00
Lithium	6.94	0.53
Fluorine (liquid)	19.00	1.14

[Hint: Assume the atoms to be 'tightly packed' in a solid or liquid phase, and use the known value of Avogadro's number. You should, however, not take the actual numbers you obtain for various atomic sizes too literally. Because of the crudeness of the tight packing approximation, the results only indicate that atomic sizes are in the range of a few Å].

Answer 13.14:

Substance	Radius (Å)
Carbon (diamond)	1.29
Gold	1.59
Nitrogen (liquid)	1.77
Lithium	1.73
Fluorine (liquid)	1.88

Atomic mass of a substance = M and Density of the substance = ρ

Avogadro's number = $N = 6.023 \times 10^{23}$

Volume of N number of molecules = $\frac{4}{3} \pi r^3 N$... (i)

Volume of one mole of a substance = $\frac{M}{\rho}$... (ii)

$$\frac{4}{3} \pi r^3 N = \frac{M}{\rho} \Rightarrow r = \sqrt[3]{\frac{3M}{4\pi\rho N}}$$

For gold: $M = 197.00 \times 10^{-3} \text{ kg}$ $\rho = 19.32 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 197 \times 10^{-3}}{4\pi \times 19.32 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.77 \text{ Å}$$

Hence, the radius of a liquid nitrogen atom is 1.77 Å.

For lithium: $M = 6.94 \times 10^{-3} \text{ kg}$ $\rho = 0.53 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 6.94 \times 10^{-3}}{4\pi \times 0.53 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.73 \text{ Å}$$

Hence, the radius of a lithium atom is 1.73 Å.

For liquid fluorine: $M = 19.00 \times 10^{-3} \text{ kg}$ and $\rho = 1.14 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 19 \times 10^{-3}}{4\pi \times 1.14 \times 10^3 \times 6.023 \times 10^{23}} \right)^{\frac{1}{3}} = 1.88 \text{ Å}$$

Hence, the radius of a liquid fluorine atom is 1.88 Å.

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Mastering Physics Chapter 13 Solutions

KJ Lindholm-Leary



Mastering Physics Chapter 13 Solutions:

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physics in a readable and interesting manner that is accessible and clear and to teach readers by anticipating their needs and difficulties without oversimplifying Physics is a description of reality and thus each topic begins with concrete observations and experiences that readers can directly relate to We then move on to the generalizations and more formal treatment of the topic Not only does this make the material more interesting and easier to understand but it is closer to the way physics is actually practiced

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Market Description This book is written for readers interested in learning the basics of physics

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