

Chapter 9

Hydrocarbons

Question 9.23:

Out of benzene, m-dinitrobenzene and toluene which will undergo nitration most easily and why?

Answer 9.23

The ease of nitration depends on the presence of electron density on the compound to form nitrates. Nitration reactions are examples of electrophilic substitution reactions where an electron-rich species is attacked by a nitronium ion (NO_2^+).

Now, CH_3 - group is electron donating and NO_2 - is electron withdrawing. Therefore, toluene will have the maximum electron density among the three compounds followed by benzene.

On the other hand, m-Dinitrobenzene will have the least electron density. Hence, it will undergo nitration with difficulty. Hence, the increasing order of nitration is as follows:



Question 9.24:

Suggest the name of a Lewis acid other than anhydrous aluminium chloride which can be used during ethylation of benzene.

Answer 9.24

The ethylation reaction of benzene involves the addition of an ethyl group on the benzene ring. Such a reaction is called a Friedel-Craft alkylation reaction. This reaction takes place in the presence of a Lewis acid.

Any Lewis acid like anhydrous FeCl_3 , SnCl_4 , BF_3 , etc. can be used during the ethylation of benzene.

Question 9.25:

Why is Wurtz reaction not preferred for the preparation of alkanes containing odd number of carbon atoms? Illustrate your answer by taking one example.

Answer 9.25

Wurtz reaction is limited for the synthesis of symmetrical alkanes (alkanes with an even number of carbon atoms) In the reaction, two similar alkyl halides are taken as reactants and an alkane, containing double the number of carbon atoms, are formed.

Mastering Chemistry Chapter 9 Answers

Elizabeth Kean, Catherine Middlecamp



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Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science CBSE Class 12
Mastering Organic Reactions MCQ With Answers Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students, educators and researchers by focusing on reaction mechanisms, practical applications and problem solving techniques. Organic chemistry is not just about memorizing equations and formulas; it is about understanding how molecules interact, change and influence each other under different conditions. With that in mind, this book emphasizes the logic and patterns behind organic reactions, making it easier for readers to apply concepts across a variety of scenarios. Each chapter of this book builds upon foundational knowledge, ensuring a progressive learning experience. From nucleophilic substitutions to pericyclic reactions and from oxidation-reduction mechanisms to named reactions, we cover both fundamental and advanced topics to cater to students at all levels. Real-world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals, biochemistry, agriculture and environmental science. This approach bridges the gap between theory and practical applications, helping readers appreciate the relevance of organic chemistry in daily life.

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