

### Chem 113 Tutorial 3 Calorimetry

1. Track through units of conversion and use equation  $C = q / \Delta T$

$\Delta H$ =heat combustion of benzoic acid (-26.42 kJ/g)

$s$ =specific heat capacity (5.02kJ/°C)

$m$ =mass (1.785g)

Initial temp = 24.62 °C

- Heat of combustion of benzoic acid is -26.42kJ/g \* 1.785g of benzoic acid = -47.1597 kJ
  - $C = \frac{q}{\Delta T}$  is rearranged to  $\Delta T = \frac{q}{C}$
  - $\Delta T = \frac{-47.1597 \text{ kJ}}{5.02 \text{ kJ/}^\circ\text{C}}$
  - $\Delta T = -9.394^\circ\text{C}$
  - $24.62^\circ\text{C} + 9.394^\circ\text{C} = 34.01^\circ\text{C}$
  - The final temperature of the bomb calorimeter is ~34.0°C.
- 

2. Use equation  $\Delta w = \text{Pressure} \times \Delta \text{Volume}$  and convert atm to Pascals (needed for appropriate units) and Litres to m<sup>3</sup>

- 1atm = 101325 Pascals so 75 atm = 7599375 Pascals  
24L is equivalent to 0.024m<sup>3</sup>
- Pressure = force/m<sup>2</sup>  
Pa is equivalent to N / m<sup>2</sup> so 7599375 pa = 7599375 N / m<sup>2</sup>
- $\Delta w = \text{Pressure} \times \Delta \text{Volume}$   
 $\Delta w = (7599375 \text{ N / m}^2) \times (0.024\text{m}^3)$
- $\Delta w = 1.8 \times 10^5 \text{ N}\cdot\text{m}$ , N\*m is equal to J

# Mastering Chemistry Chem 113 Exercise Answers

**Lingsheng Yao**



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## **Table of Contents Mastering Chemistry Chem 113 Exercise Answers**

1. Understanding the eBook Mastering Chemistry Chem 113 Exercise Answers
  - The Rise of Digital Reading Mastering Chemistry Chem 113 Exercise Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Chem 113 Exercise Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mastering Chemistry Chem 113 Exercise Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Chem 113 Exercise Answers
  - Personalized Recommendations
  - Mastering Chemistry Chem 113 Exercise Answers User Reviews and Ratings
  - Mastering Chemistry Chem 113 Exercise Answers and Bestseller Lists
5. Accessing Mastering Chemistry Chem 113 Exercise Answers Free and Paid eBooks
  - Mastering Chemistry Chem 113 Exercise Answers Public Domain eBooks
  - Mastering Chemistry Chem 113 Exercise Answers eBook Subscription Services
  - Mastering Chemistry Chem 113 Exercise Answers Budget-Friendly Options

6. Navigating Mastering Chemistry Chem 113 Exercise Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Mastering Chemistry Chem 113 Exercise Answers Compatibility with Devices
  - Mastering Chemistry Chem 113 Exercise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mastering Chemistry Chem 113 Exercise Answers
  - Highlighting and Note-Taking Mastering Chemistry Chem 113 Exercise Answers
  - Interactive Elements Mastering Chemistry Chem 113 Exercise Answers
8. Staying Engaged with Mastering Chemistry Chem 113 Exercise Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mastering Chemistry Chem 113 Exercise Answers
9. Balancing eBooks and Physical Books Mastering Chemistry Chem 113 Exercise Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mastering Chemistry Chem 113 Exercise Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Chem 113 Exercise Answers
  - Setting Reading Goals Mastering Chemistry Chem 113 Exercise Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Chem 113 Exercise Answers
  - Fact-Checking eBook Content of Mastering Chemistry Chem 113 Exercise Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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