

## Determination of $K_a$ : Titration of a Weak Acid Lab Report

11/2/22

### **Introduction:**

The goal of this lab is to determine the acid dissociation constant,  $K_a$ , by using Logger Pro and LabPro to follow the pH changes during the acid-base titration of the weak acid,  $\text{CH}_3\text{COOH}$  (acetic acid). The  $K_a$  values will come from the pH of the original solution and the pH of the half-neutralization point. We will also be finding the  $[\text{HA}]$  concentration using the equivalence volume by looking at our titration curve.

Equilibrium is represented by the acid dissociation constant ( $K_a$ ):

$K_a =$

The equation of the reaction is as follows:



### **Procedures:**

- 1) Add 20.0 mL of weak acid solution into 100 mL beaker

# Determination Of Ka Lab Report Answers

**Lingsheng Yao**



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