Balancing Equations & Types of Reactions WS 1

Practice Problems

In your notebook, solve the following problems.

NOTE: A skeleton equation is an unbalanced equation.

SECTION 10.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.

2H2(9) + 02(8) = 2H20(1)

2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.

Fe(s) + S(s) -> Fes(s)

 Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.

Mg CO3 -> MgO + CO2 (0

4. Write a balanced equation for the production of HCl gas from its elements.

H2(g) + C12(g) -> 2HC1(g)

5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.

 $2HCl(\alpha q) + CaCO_1(x) \rightarrow CO_2(g) + CaCl_2(\alpha q) + H_2O(f)$

Hydrichinic acid and Calcium react to annound and water.

Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the
equation correctly balanced? Explain.

 $2Ag(s) + S(s) \rightarrow Ag_2S(s)$

2 solid shuter reacts wy I solid sulfter to create solid silver sulfide.

Describing Chemical Reactions Section Review Answer Key

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