

Determination of the Solubility-Product Constant for a Sparingly Soluble Salt

A. Preparation of a Calibration Curve

Initial $[\text{CrO}_4^{2-}]$ 0.0024M

Volume of 0.0024 M K_2CrO_4	Total volume	$[\text{CrO}_4^{2-}]$	Absorbance
1. <u>1 mL</u>	<u>50 mL</u>	<u>0.00048M</u>	<u>0.190</u>
2. <u>5 mL</u>	<u>50 mL</u>	<u>0.0024M</u>	<u>1.039</u>
3. <u>10 mL</u>	<u>50 mL</u>	<u>0.0048M</u>	<u>2.127</u>
4. <u>15 mL</u>	<u>50 mL</u>	<u>0.0072M</u>	<u>2.409</u>

Molar absorption coefficient for $[\text{CrO}_4^{2-}]$

1. _____ 2. _____ 3. _____ 4. _____

Average molar absorption coefficient _____

Standard deviation (show calculations) _____

B. Determination of the Solubility-Product Constant

Absorbance	$[\text{CrO}_4^{2-}]$	$[\text{Ag}^+]$	K_{sp} of Ag_2CrO_4
1. <u>0.603</u>	_____	_____	_____
2. <u>0.753</u>	_____	_____	_____
3. <u>1.286</u>	_____	_____	_____

Determination Of A Solubility Product Lab Answers

Colleen F. Craig, Kim N. Gunnerson



Determination Of A Solubility Product Lab Answers:

Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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