

Section 5.4: Hess's Law

Tutorial 1 Practice, page 317

1. (a) Solution:

Step 1: Label the equations.

The desired equation is: $2 \text{N}_2(\text{g}) + 6 \text{H}_2\text{O}(\text{g}) \rightarrow 3 \text{O}_2(\text{g}) + 4 \text{NH}_3(\text{g})$



Step 2: Identify the equations that have reactants or products on the same side of the arrow as in the desired equation. Reverse (1) and (2) so that reactants and products are on the same side as in the desired equation. Reverse the sign of ΔH .



Step 3: Check the coefficients of all entities in the equations to see if any must be adjusted by a factor so as to be equal to the coefficients in the desired equation.

Equation (3) has $\frac{1}{2}$ mol of nitrogen gas on the reactant side, while the desired equation has

2 mol. Multiply equation (3) and its ΔH by the factor 4:



Equation (4) has 2 mol of water on the reactant side, while the desired equation has 6 mol.

Multiply equation (4) and its ΔH by the factor 3:

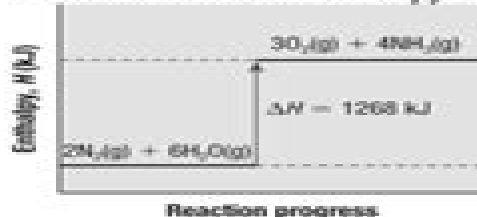


Step 4: Add equations (5) and (6), and their changes in enthalpies.



Statement: The ΔH for the reaction is 1268 kJ.

(b) Draw the reaction coordinates with the x -axis labelled "Reaction progress" and the y -axis labelled "Enthalpy, H (kJ)." Since ΔH is positive, the products will have greater enthalpy than the reactants. Point the arrow from reactants upward to products. Write the reactants on the lower left-hand side of the graph and the products on the upper right-hand side of the graph. Include the ΔH of the reaction. An enthalpy diagram of the reaction is:



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Padhraic Smyth



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