8.2 - THE CHEMICAL EARTH

1 - The living and non-living components of the Earth contain mixtures

1.1 Construct word and balanced formulae equations of chemical reactions as they are encountered

To make word equations, write each chemical formula, in words. Then, add in pluses and arrows.

magnesium + oxygen → magnesium oxide

 To create balanced formulae equations, write down chemical symbols (see 1.7), their state of matter and balance the equation.

Magnesium + Oxygen → Magnesium Oxide	The original word equation.
$Mg + O_2 \rightarrow MgO$	There are not enough Oxygen atoms on the right.
$Mg + O_2 \rightarrow 2MgO$	Double the whole molecule, not just magnesium. There are not enough magnesium atoms now.
2Mg + O ₂ → 2MgO	Add a 2 in front of magnesium on the left. Now, LHS = RHS
2Mg _{(cl} + O _{2 (gl} → 2MgO _{(cl}	Place the states of matter for each substance.

- Combustion Fuel + Oxygen → Carbon dioxide + Water, example: CH₄ + 2O₇ → CO₇ + 2H₂O
- Corrosion Metal + Oxygen → Metal Oxide, example: 2Cu + O₂ → 2CuO
- Respiration Glucose + Oxygen → Carbon Dioxide + Water, example: C₆H₁₂O₆ + 6O₂ → 6CO₂ + 6H₂O
- Precipitation Soluble salt + Soluble salt → Insoluble salt + Soluble salt
 - Example: CuSO₄ + 2NaOH → Cu(OH)₂ + Na₂SO₄ (Cu(OH)₂ is insoluble)
- Neutralisation AcId + Base → Salt + Water, example: HCl + NaOH → NaCl + H₂O
- Acids on metals Metal + Acid → Salt + Hydrogen, example: Zn + H₂SO₄ → ZnSO₄ + H₂
- Acid on metal oxide Metal Oxide + Acid → Salt + Water, example: MeO +2HCl → MeCl₂ + H₂O
- Acids on carbonates Carbonate + Acid → Salt + Water + Carbon Dioxide
 - Example: CaCO₃ + 2HCl → CaCl₂ + H₂O + CO₂
- Decomposition Molecule → Molecule part + Molecule part, example: 2H₂O → 2H₂ + O₂

1.2 Identify the difference between elements, compounds and mixtures in terms of particle theory.

- Particle theory helps us visualise differences between elements, compounds and mixtures.
- Elements are basic building blocks, having own unique symbol and physical and chemical properties.
- Compounds are two or more particles chemically bonded.
 - A mixture is composed of a mix of elements and compounds, e.g. a solution.
 - Can be separated using simple separation techniques, does not have fixed properties.

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