

Chapter Test B

Chapter: States of Matter

Match each term with its definition, using the number of the term in parentheses that completes each statement correctly. Write the number in the space.

1. According to the kinetic-molecular theory, particles of a substance
 - a. move about with both the best and worst results.
 - b. move about with equal results.
 - c. move both the best and worst results.
 - d. do not exhibit random motion at any temperature.
2. According to the equation $KE = \frac{1}{2}mv^2$, all particles within a substance
 - a. move about with the same energy.
 - b. move about with the same speed.
 - c. have different speeds.
 - d. have different energies.
3. Liquids differ from gases mainly because
 - a. liquids have less compressibility.
 - b. the particles do not have random motion.
 - c. they possess the highest intermolecular forces.
 - d. liquids are composed of atoms, whereas gases are made of molecules.
4. When heated, a gas
 - a. expands without change in volume.
 - b. moves from a high temperature region to a low temperature region.
 - c. exhibits a change in density resulting in expansion.
 - d. moves from a high temperature region to a low temperature region.
5. When heated, the force of attraction for a liquid is
 - a. reduced, and the liquid will expand.
 - b. reduced, and the liquid will contract.
 - c. increased, and the liquid will expand.
 - d. increased.
6. A solid substance shows expansion when heated because
 - a. the particles move more rapidly.
 - b. they absorb heat energy.
 - c. the particles move more rapidly.
 - d. the particles are held together.

Assessment Chapter Test B States Of Matter

Danica G. Hays



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