

CHEM 1001

ATOMIC SPECTRA Laboratory Report

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SECTION: 10/21

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A. Purpose/Guide of Experiment:

The goals of the experiment are to understand atomic energy levels and the discrete colour lines or absorption transitions and to use spectrometer and quantitative data to infer the most common ionization frequencies of known elements.

B. Background Theory:

Energy is stored in different atomic levels (energy levels) as a result of $n=1$ specific orbital, contributing their distance from the nucleus. Ground state is the lowest amount of energy $n=1$ where the atom, molecules are represented by where the atom has an infinite amount of energy. Electrons transition to their movement of infinite and frequency energy levels. Electrons move to their energy absorption and emission. Absorption is where electrons absorb the energy used jumps to a higher energy level. Emission mechanism is the opposite - the electrons release energy and jumps to a lower energy level. Bohr's Theory states that an atom consists of a positively charged and a negatively charged nuclei quantised energy n (or orbits), specific to hydrogen atoms.

C. Results and calculations:

Figure 1: Spectrum of the hydrogen atom

Results of Hydrogen, shown in the corresponding Wavelength/Nucleus Transition:

Observed w. length of	Observed Wavelength/nm	Calculated Wavelength/ nm (n=2 to n=1)	n_{upper}	n_{lower}
656	656.3 nm	656.3 nm	2	1
486	486.1 nm	486.1 nm	3	2
434	434.0 nm	434.0 nm	4	3
410	410.2 nm	410.2 nm	5	4

Atomic Spectra Lab Report Answers

BM King



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Unveiling the Magic of Words: A Overview of "**Atomic Spectra Lab Report Answers**"

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