

Colligative Properties and Activity Coefficients Exercise 1 and 2 (continued)

Knowledge being tested is your ability to calculate the _____

Solution

For solution of Exercise 1 (continued):

What does including activity coefficients?

What is the molality of a solution prepared by dissolving 5.00 g of sodium chloride (NaCl) in 100 g of water?

$$\frac{5.00 \text{ g NaCl} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{100 \text{ g} \times \frac{1 \text{ kg}}{1000 \text{ g}}} = \frac{0.0855 \text{ mol NaCl}}{1 \text{ kg water}}$$

Solution:

$$\frac{5.00 \text{ g NaCl} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{100 \text{ g} \times \frac{1 \text{ kg}}{1000 \text{ g}}} = \frac{0.0855 \text{ mol NaCl}}{1 \text{ kg water}} = \frac{0.0855 \text{ mol NaCl} \times \frac{1 \text{ mol}}{58.44 \text{ g}}}{100 \text{ g} \times \frac{1 \text{ kg}}{1000 \text{ g}}} = \frac{0.00146 \text{ mol NaCl}}{1 \text{ kg water}}$$

Temperature-Dependent Activity Coefficients

This paper presents the solution to the following question:

When a solution of a solute is prepared, the activity of the solute is _____ (Figure 10.1, Table 10.1, the value of the activity of the pure solvent).

Colligative Properties Study Guide Answers

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