

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfate, the roman numeral tells us (2)
 - How many atoms of the anion are in the formula unit.
 - How many sulfate ions can be attached to the iron atom.
 - The charge on each Fe ion.
 - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (3)
 - Change the number of moles represented in the formula.
 - Change the charges on the other ions in the compound.
 - Change the formula so that it no longer represents the compound.
 - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{HC}_2\text{H}_3\text{O}_2$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II chloride	PbCl_2
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Hydrochloric acid	HCl
Dinitrogen pentoxide	N_2O_5
Copper II carbonate	CuCO_3
Sodium sulfate	Na_2SO_4
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Barium sulfate	BaSO_4
Nitrous acid	HNO_2
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Sulfur trioxide	SO_3
Sulfur dioxide	SO_2
Barium perchlorate	$\text{Ba}(\text{ClO}_4)_2$
Tetraphosphorus trisulfide	P_4S_3
Phosphoric acid	H_3PO_4

- Name the following compound and identify its molecular formula (2)
 - $\text{Ca}(\text{NO}_3)_2$, calcium (II) nitrate - II
 - $\text{Fe}(\text{NO}_3)_3$, iron (III) nitrate - III
 - $\text{Fe}(\text{ClO}_4)_3$, iron (III) perchlorate - III
 - N_2O_5 , dinitrogen pentoxide - III

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