

Chapter 21 -- Nuclear Chemistry

21.1 Radioactivity

- When nuclei change spontaneously, emitting energy, they are said to be **radioactive**.
- **Nuclear chemistry** is the study of nuclear reactions and their uses.
- *Isotopes* have the same number of p⁺ but different numbers of n⁰.
 - Different isotopes of the same element are distinguished by their mass numbers.
 - Different isotopes have different natural abundances.
- A **radionuclide** is a radioactive nucleus.
 - Atoms containing these nuclei are called **radioisotopes**.

Nuclear Equations

- Most nuclei are stable.
 - Radionuclides are unstable and spontaneously emit particles and/or electromagnetic radiation.
 - Example: Uranium-238 is radioactive.
 - It emits **alpha (α) particles** (helium-4 nuclei)
- When a nucleus spontaneously decomposes in this manner, we say it has decayed (*radioactive decay*).
- In nuclear equations, the total number of nucleons is conserved.
 - We can represent the uranium-238 decay by the following nuclear equation:



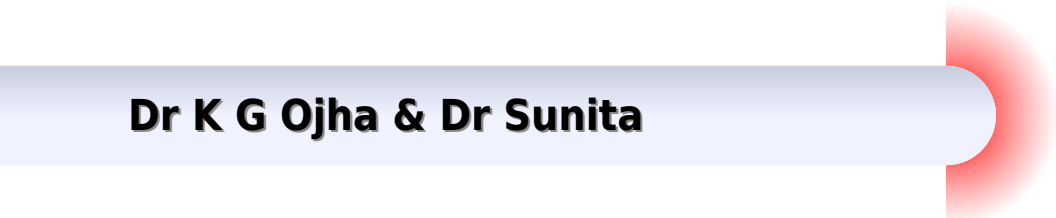
- The total number of protons and neutrons before a nuclear reaction must be the same as the total number of nucleons after the reaction.

Types of Radioactive Decay

- There are three types of radiation which we will consider:
 - α-Radiation is the loss of ${}_2^4\text{He}$ (**alpha particles**) from the nucleus.
 - β-Radiation is the loss of an electron from the nucleus. These high-speed electrons are called **beta particles**.
 - γ-Radiation is the loss of high-energy photons from the nucleus.
- Nucleons can undergo two other types of decay:
 - **positron emission**
 - A positron is a particle with the same mass as an electron but with the opposite sign.
 - **electron capture**
 - The nucleus captures an electron from the electron cloud surrounding the nucleus.
- Representations:
 - In nuclear chemistry, to ensure the conservation of nucleons we write all particles with their atomic and mass numbers: ${}_2^4\text{He}$ and ${}_{-1}^0\text{e}$ represent α- and β- radiation.
- Nucleons can undergo decay:
 - (β-emission) ${}_0^1\text{n} \rightarrow {}_1^1\text{p} + {}_{-1}^0\text{e}$
 - (positron annihilation) ${}_1^0\text{e} + {}_{-1}^0\text{e} \rightarrow 2{}_0^0\gamma$
 - (positron or β⁺-emission) ${}_1^1\text{p} \rightarrow {}_0^1\text{n} + {}_1^0\text{e}$
 - (electron capture) ${}_1^1\text{p} + {}_{-1}^0\text{e} \rightarrow {}_0^1\text{n}$

Chapter 21 Nuclear Chemistry Section 3 Answers

Dr K G Ojha & Dr Sunita



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Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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