

Chapter 6 Thermochemistry

Homework Problems

- The enthalpy change for the fermentation (without oxygen, O_2 gas) reaction of glucose to alcohol and carbon dioxide is -67 kJ .
 - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
 - Is the reaction exothermic or endothermic?
 - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is $0.71 \text{ J/}^\circ\text{C}\cdot\text{g}$. Calculate the energy required to:
 - Raise the temperature of 1.0 mol graphite by 1.0°C .
 - Raise the temperature of 850 g graphite by 150°C .
 - Raise the temperature of 75 kg graphite from 294 K to 348 K .
- In a coffee-cup calorimeter, 100.0 mL of 1.0 M NaOH and 100.0 mL of 1.0 M HCl are mixed. Both solutions were originally at 24.6°C . After the reaction, the final temperature is 31.3°C . Assume that all the solutions have a density of 1.0 g/mL and a specific heat capacity of $4.18 \text{ J/g}\cdot^\circ\text{C}$. Calculate the enthalpy change for the neutralization of HCl by NaOH . Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ($\text{C}_{15}\text{H}_{26}\text{O}$) has an energy of combustion of -5903.6 kJ/mol . A sample of camphor with a mass of 0.1204 g is burned in a bomb calorimeter, the temperature increases by 2.28°C . Calculate the heat capacity of the calorimeter.

A 0.1964 g sample of quinone ($\text{C}_6\text{H}_4\text{O}_2$) is burned in a bomb calorimeter that has a heat capacity of $1.56 \text{ kJ/}^\circ\text{C}$. The temperature of the calorimeter increases by 3.2°C . Calculate the energy of combustion of quinone per gram and per mole.

Chapter 6 Thermochemistry Homework Problems

L Cohen



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