

CHAPTER 7 REVIEW*Chemical Formulas and Chemical Compounds***SECTION 7-1****SHORT ANSWER** Answer the following questions in the space provided.

1. C In a Stock name such as iron(III) sulfate, the roman numeral tells us _____
(a) how many atoms of Fe are in one formula unit
(b) how many sulfate ions can be attached to the iron atom
(c) the charge on each Fe ion
(d) the total positive charge of the formula unit
2. C The result of changing a subscript in a correctly written chemical formula is to _____
(a) change the number of moles represented by the formula
(b) change the charges on the other ions in the compound
(c) change the formula so that it no longer represents the compound it previously represented
(d) have no effect on the formula
3. The explosive TNT has the molecular formula $C_7H_5(NO_2)_3$.
4 elements a. How many elements make up this compound?
6 oxygen atoms b. How many oxygen atoms are present in one molecule of $C_7H_5(NO_2)_3$?
21 atoms c. How many atoms in total are present in one molecule of $C_7H_5(NO_2)_3$?
 4.2×10^{24} atoms d. How many atoms are present in a sample of 2×10^{23} molecules of $C_7H_5(NO_2)_3$?
4. How many atoms are present in each of these formula units?
11 atoms a. $Ca(HCO_3)_2$
45 atoms b. $C_{12}H_{22}O_{11}$
10 atoms c. $Fe(ClO_2)_3$
9 atoms d. $Fe(ClO_3)_2$
5. N_2O_5 a. What is the formula for the compound dinitrogen pentoxide?
carbon(IV) sulfide b. What is the Stock name for the covalent compound CS_2 ?
 H_2SO_3 c. What is the formula for sulfurous acid?
phosphoric acid d. What is the name for the acid H_3PO_4 ?

Chapter 9 Chemical Names And Formulas Section Review Answers

Dennis D. Staley, Antony C. Wilbraham



Chapter 9 Chemical Names And Formulas Section Review Answers:

Understanding Chemistry Fred M. Dewey, 1994 Essential Concepts of Chemistry Sharon Sherman, Alan Sherman, 1999
Designed especially for students who have little or no background in chemistry or mathematics Essential Concepts of Chemistry makes complex concepts understandable This text provides an inexpensive one color alternative for introductory chemistry courses and emphasizes everyday applications of chemistry Chemistry Henry Dorin, 1987 Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Chemistry George Tyler Miller, 1987 Essentials of Chemistry Dennis D. Staley, Antony C. Wilbraham, 1984 **Prentice Hall Chemistry**, 2000 *Foundations of College Chemistry* Morris Hein, Susan Arena, Cary Willard, 2023 *Foundations of College Chemistry* 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and applications of radioactivity *Holt Chemistry* Ralph Thomas Myers, 2004 *Introductory Chemistry* Victor S. Krimsley, Darold E. Skeritt, Beverly B. Harrison, 1986 This pedagogically rich text has all the necessary features to hook introductory students and keep them interested and successful in preparatory chemistry This book carefully guides beginning students through the fundamental principals and calculations required for their subsequent success in either the general chemistry course or the short organic and biochemistry course Krimsley while dedicated to conceptual understanding and skill building presents a solid book that provides students with complete explanations on every point and helps them work through the material methodically with many examples and hints His approach is cohesive and coherent always patient and interesting Krimsley introduces all topics through an example students are already familiar with He continually reminds them of objectives and provides many opportunities to practice and then check their answers The text begins with a study of atomic and molecular structure prior to treating the various categories of chemical reactions The organization is designed to get students into chemistry quickly yet methodically The classification of matter begins in Chapter 2 before the chapter on measurement and the coverage of bonding appears in Chapter 8 The elementary concepts of chemistry are presented with an emphasis on mathematical calculations and problem solving strategies *Foundations of College Chemistry* Morris Hein, 1991 Features strong problem solving and carefully constructed problem sets **Chemistry** James E. Brady, Fred

Senese,2004-02-04 Publisher Description *Marketing Information Guide* ,1963 **Machinery** ,1907 **Machinery**
Lester Gray French,1907 **Basic Chemistry** Steven S. Zumdahl,1999 For a full description see catalog entry for Zumdahl
Introductory Chemistry A Foundation 4 e Foundations of College Chemistry, Alternate Morris Hein,Susan
Arena,2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For
over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and
logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked
out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of
chemistry and relates chemistry to things health professionals experience on a regular basis *Railroad Gazette* ,1894
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