

General Chemistry I

Practice Test with Answer Key

Chemical Reaction

Test I: True or False. Write **True** if the statement is correct and write **False** if the statement is wrong.

- _____ 1) chemical reactions can be easily analyzed if they are written through chemical equations
- _____ 2) Reactants are substances produced after a chemical reaction.
- _____ 3) Chemical Equation is the shorthand for describing the course of chemical reaction
- _____ 4) Chemical Reaction is a process in which one or more substances are changed into new substance with the same physical and chemical properties
- _____ 5) (g) is an indication that the substance is in gas form
- _____ 6) Main key to balance an equation is to write the correct subscript before the chemical symbol
- _____ 7) Converting the word equation into a chemical equation is one of the pre-step in balancing
- _____ 8)  means heat is applied during the reaction
- _____ 9) (aq) means aqueous, substance is dissolved in ammonia
- _____ 10) The water we drink every day, the sugar found in bread, the carbon dioxide gas produced from burning a piece of paper are examples produced as the result of chemical reactions.

Test II: Multiple Choices. Choose the **best answer** among the choices. Write the letter of the correct answer before the number.

- _____ 1) Substances involved before a chemical reaction
a. *product* b. *reactant* c. *compound* d. *elements*
- _____ 2) The end or final materials in a chemical reaction is known as
a. *Reactant* b. *product* c. *substance* d. *quotient*
- _____ 3) What is the meaning of + used in a chemical equation?
a. *yields* b. *reacts with* c. *heat is applied* d. *gas*
- _____ 4) In the reaction $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$, oxygen gas is
a. *subscript* b. *product* c. *reactant* d. *coefficient*
- _____ 5) Which equation is correctly balanced?
a. $\text{C} + \text{O}_2 \rightarrow 2 \text{CO}$ b. $\text{C} + \text{O}_2 \rightarrow \text{CO}$ c. $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ d. $2 \text{C} + \text{O} \rightarrow 2 \text{CO}$
- _____ 6) In the equation $2 \text{C}_2\text{H}_2(\text{g}) + 5 \text{O}_2 \rightarrow 4 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$, the coefficient for carbon dioxide is
a. 2 b. 5 c. (g) d. 4
- _____ 7) All of the following are chemical reactions EXCEPT

- a. $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{g})$
- b. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{aq})$
- c. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- d. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- e. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- f. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- g. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- h. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- i. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- j. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- k. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- l. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- m. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- n. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- o. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- p. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- q. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- r. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- s. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- t. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- u. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- v. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- w. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- x. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- y. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- z. $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$

Test III: Multiple Choices. Choose the **best answer** among the choices. Write the letter of the correct answer before the number.

- _____ 1) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{aq})$
- _____ 2) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{g})$
- _____ 3) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- _____ 4) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- _____ 5) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- _____ 6) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- _____ 7) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- _____ 8) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$
- _____ 9) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{aq}) + \text{NaCl}(\text{g})$
- _____ 10) $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{g}) + \text{NaCl}(\text{aq})$

Chemical Reaction Test And Answer

Joacim Rocklöv



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