Chemistry: Balancing Chemical Equations

<u>Directions</u>: First, balance each of the chemical equations below. Then, classify each reaction as **synthesis**, **decomposition**, **single-replacement**, or **double-replacement**. To earn full credit, write the words out when classifying.

Balance the equation...

- _2_Sb + _3_Cl₂ → _2_SbCl₃
- 2. Mg + ___O₂ → __2_MgO
- CaCl₂ → ___ Ca + ___ Cl₂
- 2 NaClO₂ → 2 NaCl + 3 O₂
- Fe + 2 HCl → FeCl₂ + H₂
- 6. CuO + H₂ → Cu + H₂O
- 7. 2 Al + 3 $H_2SO_4 \rightarrow$ Al₂(SO₄)₁ + 3 H_2
- MgBr₂ + ___ Cl₃ → ___ MgCl₂ + ___ Br₂
- 9. ____SnO₂ + __2 C → ____Sn + __2 CO
- 10. $Pb(NO_3)_2 + ___ H_2S \rightarrow ___ PbS + __2 HNO_3$
- 11. __2_HgO → __2_Hg + ____O;
- 12. __2_KClO₃ → __2_KCl + __3_O₂
- 13. $N_2 + 3 H_2 \rightarrow 2 NH_3$
- 14. __2 NaBr + ___ Cl₂ → __2 NaCl + ___ Br₂
- 15. ____ Zn + __2_ AgNO₃ → ____ Zn(NO₃)₂ + __2_ Ag
- 16. ____ Sn + __2 Cl₂ → ____ SnCl₄
- 17. ____ Ba(OH)₂ → ____ BaO + ____ H₂O

...and classify it.

synthesis

synthesis

decomposition

decomposition

single replacement

single replacement

single replacement

single replacement

single replacement

double replacement

decomposition

decompostion

synthesis

single replacement

single replacement

synthesis

decomposition

Chemical Calculations And Equations Answer Key

John McMurry, Robert C. Fay

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