

Chemistry Worksheets Class 11 on Chapter 8 Redox Reactions - Set 1

Q-1: Which of the following statements about redox reactions is correct?

- a) A nonmetal with a higher reduction potential is a more potent reducing agent.
- b) A metal with a lower oxidation potential is a more potent oxidising agent.
- c) The oxygen oxidation state in O_3 is -1.
- d) All of the above

Q-2: Which of the following can behave as a strong reducing agent?

- a) $HClO_4$
- b) $SnCl_2$
- c) SO_2
- d) All of the above

Q-3: The pair of compounds in which both metals are in the highest oxidation state is

- a) $FeSO_4$, CuS_2
- b) CrO_2Cl_2 , MnO_4^-
- c) TiO_2 , MnO_2
- d) $[Co(CN)_6]^{3-}$, MnO_2

Q-4: The number of electrons involved when 1 mole of H_2O_2 decomposes as

- $H_2O_2 \rightarrow H_2O + O_2$
- a) 1
 - b) 6
 - c) 8
 - d) 4

Q-5: Zinc iodide is formed during the reaction of zinc and iodine. Which of these is being oxidised?

- a) Zinc ions
- b) Zinc atom
- c) Iodide ions
- d) Iodine

Q-6: Can a solution of 1M copper sulphate be stored in a vessel made of nickel metal?

Given that $E^\circ(Ni^{2+}/Ni) = -0.25$ volt and $E^\circ(Cu^{2+}/Cu) = +0.34$ volt.

Q-7: Using the electronic concept, identify the oxidant and reductant in the following redox reactions.

- i) $Zn + 2H^+ \rightarrow Zn^{2+} + H_2$
- ii) $BrO_3^- + F_2 + 2OH^- \rightarrow BrO_4^- + 2F^- + H_2O$

Chemistry Chapter 11 Chemical Reactions Packet Answers

Albert Kaskel



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Chemical Misconceptions Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

Books in Print Supplement, 2002 [Dynamics of Molecule Surface Interaction](#) Gert Due Billing, 2000-01-03 Chemical reactions at surfaces do not follow the same reaction dynamics as atoms in the gas phase The changes from ideal interactions depend on the electronic structure and the spatial and geometric shape of the surface The dynamics of chemical reactions at surfaces have all the complexity of gas phase molecules reaction dynamics plus those associated with the additional phenomena due to the presence of a solid surface

Advanced Molecular Dynamics and Chemical Kinetics Gert Due Billing, Kurt V. Mikkelsen, 1997-04-11 A comprehensive in depth presentation of theoretical underpinnings and mathematical techniques This is the first book of its kind to combine all the theories of molecular reaction dynamics and chemical kinetics in a single source It provides a sophisticated treatment of the material that functions both as a professional reference and a high level text for PhD and postdoctoral researchers Advanced Molecular Dynamics and Chemical Kinetics offers exceptional in depth coverage and includes a complete discussion of the theoretical as well as mathematical presentation of techniques It features relevant exercises as well as comprehensive coverage of Second Quantization Semiclassical Theory Quantum Theory of Reaction Rates Feynman Path Integrals Wavepacket Propagation and Grid Methods Photodissociation Molecular Properties of Solvated Molecules Quantum Model for Electron Transfer Electron Transfer Coupling Elements Proton Transfer Reactions in Solution This is the ideal reference for seasoned professionals in molecular reaction dynamics as well as for younger researchers who may want to enter the field or simply wish to learn more about it Also available Introduction to Molecular Dynamics and Chemical Kinetics Gert D Billing and Kurt V Mikkelsen

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