

Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
 - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
 - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
 - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} \cdot \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{1 \text{ mol CaO}} \cdot \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$. Remember the ratios can be reduced as well!!!
 - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} \cdot \frac{1 \text{ mol H}_2}{1 \text{ mol H}_2} \cdot \frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol MgF}_2}{2 \text{ mol HF}} \cdot \frac{1 \text{ mol H}_2}{1 \text{ mol H}_2} \cdot \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol H}_2}{1 \text{ mol MgF}_2}$
- $\frac{165.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{165.99 \text{ g Na}_2\text{CO}_3} \cdot \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \cdot \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} \cdot \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} \cdot \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \cdot \frac{108.99 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} \cdot \frac{1 \text{ mol CaCO}_3}{108.99 \text{ g CaCO}_3}$
- Molar Mass
 - Definition: the mass of one mole of a substance
 - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen $\rightarrow$ Water
 - 5.0 mol H₂
 - 2.5 mol O₂
- Ethane combustion reaction
 - 15.8 mol O₂
 - 9.00 mol CO₂ & 13.5 mol H₂O
- 575 g Na & 886 g Cl₂
- Iron production reaction
 - 75.1 mol CO
 - 50.1 mol Fe & 75.1 mol CO₂
- 87,390 g CO or $8.739 \times 10^5 \text{ g CO}$ & 12,610 g H₂ or $1.261 \times 10^4 \text{ g H}_2$
- Lightning reaction of N₂ & O₂
 - 1,100 g NO₂ or $1.10 \times 10^3 \text{ g NO}_2$
 720. g NO or $7.20 \times 10^2 \text{ g NO}$
- CO₂ removal in spacecraft
 - 42.04 mol NaOH
 - 21.02 mol Na₂CO₃ & 21.02 mol H₂O
- Double replacement AgNO₃ & NaBr
 - 463 g NaBr
 - 845 g AgBr
- Soda-acid fire extinguisher
 - 3.058 mol NaHCO₃
 - 3.058 mol CO₂, 1.529 mol Na₂SO₄, 3.058 mol H₂O
- Sulfuric acid-sodium hydroxide neutralization reaction
 - H₂SO₄ + 2 NaOH $\rightarrow$ Na₂SO₄ + 2 H₂O
 - 37 g H₂SO₄
 - 53 g Na₂SO₄ & 14 g H₂O
- Copper & silver nitrate reaction
 - 0.0104 mol Cu(NO₃)₂
 - 0.0104 mol Cu & 0.0209 mol AgNO₃
- Aspirin production
 - 13.5 kg C₆H₅O₄
 - 7.66 kg C₆H₅O₃
 - 4.29 L HC₂H₃O₂
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

Chemistry Chapter 9 Stoichiometry Answers

John M. DeKorte



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Student Solutions Manual, Chemistry, Principles & Practice, Second Edition John M. DeKorte, 1997

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