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SECTION A

1. Complete the following equations for nuclear reactions.



2. The first three successive ionization energies of element Y are 549, 1064 and 4225 kJmol⁻¹.

- Explain the trend in the variation of the ionization energies of Y.
- State the group in the Periodic Table to which Y belongs.

3. Complete the equation below and write the mechanism for the reaction.



4a) Write

- equation for the ionization of benzoic acid in water.
- the expression for the ionization constant, K_a of benzoic acid.

b) Calculate the pH of a 0.1M solution of benzoic acid.
($K_a = 6.3 \times 10^{-5} \text{ mol dm}^{-3}$).

5a) 10cm³ of a hydrocarbon P, (C₈H₁₈) was exploded in 90cm³ of oxygen. On cooling to room temperature, the residual gases occupied 70cm³. When the residual gases were passed through potassium hydroxide solution, the volume reduced to 40cm³.

- Write equation for the reaction between P and oxygen.
- Determine the molecular formula of P.
- Write equations to show how P can be prepared from an alcohol.

6. When compound W was steam distilled at 86°C and 760mmHg pressure, the distillate contained 85% of water by mass. Calculate the relative molecular mass of W.

(The vapour pressure of water at 86°C is 740 mmHg)

Chemistry Paper 1 06

Mike Jess

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