PHENOLIC GROUP (ArOH) 5.2

Phenols are colourless but attain colour due to oxidation. Moreover, they are acidic and dissolve in 5% NaOH. Phenols are generally highly reactive in electrophilic aromatic substitution. On bromination (bromine water) of phenol, 2, 4, 6 - tribromophenol is obtained.

- ((i) Litmus paper test
- (iii) NaHCO3 test
- (iiii)

- (iv) Bromine water test (v) Azodye test
- (vi) Liebermann nitroso reaction

Dissolve 0.2 g of the substance in 2 mL conc. H2SO4 and add a few crystals of NaNO2 - a blue colouration develops immediately which becomes red on dilution with water. Add NaOH in excess. The colour changes to green, blue or violet. Among dihydroxyphenols only resorcinol gives a satisfactory positive test.

Chemical reaction

2NaNO₂ + H₂SO₄
$$\longrightarrow$$
 Na₂SO₄ + 2HNO₂

OH

OH

NO

OH

NOH

NOH

H₂SO₄

H₂SO₄

H₃O

Acid sulphate of indophenol
(blue colouration)

Na

NaOH

Sodium salt of indophenol
(Deep blue colouration)

Only those phenols possessing a free para position respond to this test.

Chemistry Test Functional Groups Chapter 23

Ralph L. Shriner

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