Chapter 1 - Kinetics - Practice Test #1

The answers to this test are posted online

Name -

Consider the following reaction mechanism:

step 1: M + X → MX step 2: MX + A → D + X

The chemical species MX is a(n)

- A catalyst C final product

 B inhibitor B D reaction intermediate
- Consider the following reaction: 2N₂O_{Ng} → 4NO_{Ng} + O_{Ng}
 At a certain temperature the rate of decomposition of 2N₂O_{Ng} is 2.5 x 10⁻⁶ mol/s. The rate of formation of NO₂ is

A. 1.0 x 10⁻⁵ mol/s C. 2.5 x 10⁻⁶ mol/s B. 1.3 x 10⁻⁶ mol/s D. 5.0 x 10⁻⁶ mol/s

 Which of the following factors affect the rates of both homogeneous and heterogeneous reactions.

I nature of the reactants
II presence of a catalyst
III temperature of system
IV concentration of reactants

A. I and IV only
B. II and III only

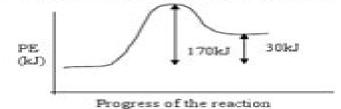
C II, III, and IV only
D I, II, III, and IV

4. Which of the following equations represents an endothermic reaction?

A N₂O_{4(g)} + 59 kJ → 2NO_{2(g)}

- B. $2H_{340} + O_{340} \rightarrow 2H_{3}O_{(3)} + 572 \text{ kJ}$
- C. $2BrCl_{100}$ -29.3 kJ $\rightarrow Br_{200}$ + Cl_{200} D. $2H_{200}$ + O_{200} \rightarrow $2H_{2}O_{00}$ ΔH = -572 kJ
- Consider the potential energy diagram. The activation energy for the reverse reaction is

A 30 kJ B 140 kJ



C 170 kJ

D. 200 kJ

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create works of wide ranging color and complexity an unorthodox implement a pair of scissors was the tool matisse used to transform paint and paper into a world of plants animals figures and shapes