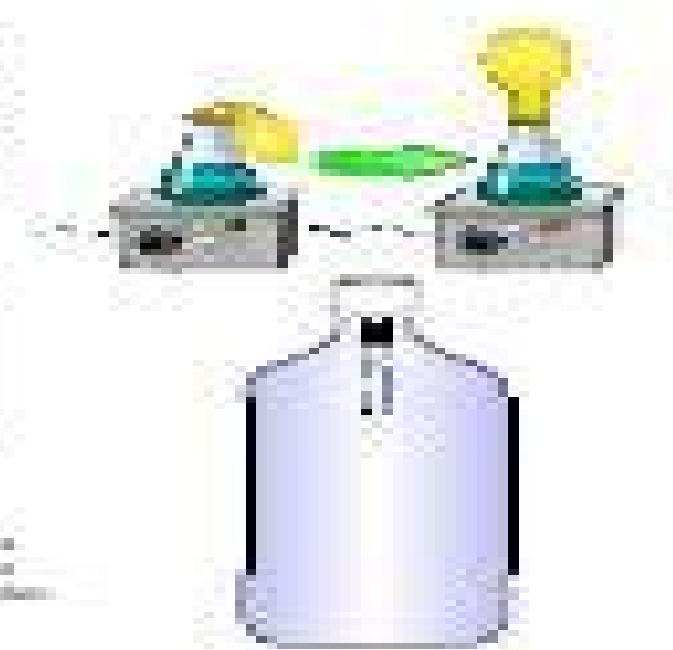


Using the Combined Gas Law

We believe in gaseous molecules only fixed and the atmosphere is a free place, the balloon floats up as it heats up. This happens even though we still think it's just the balloon. As the temperature of the gas increases, the gas expands. This is what happens when you heat up a gas. The gas expands because the molecules are moving faster. This is what happens when you heat up a gas. The gas expands because the molecules are moving faster. This is what happens when you heat up a gas.



The Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Where:

P_1 = Initial pressure

V_1 = Initial volume

T_1 = Initial temperature (K)

P_2 = Final pressure

V_2 = Final volume

T_2 = Final temperature (K)

This equation can be used to find the final pressure, volume, or temperature of a gas if you know the initial pressure, volume, and temperature. It can also be used to find the final pressure, volume, or temperature of a gas if you know the initial pressure, volume, and temperature.

Example Problem

A gas with a volume of 100 mL at 100 kPa and 300 K is heated to 400 K and the pressure is increased to 150 kPa. What is the final volume?

$$P_1 = 100 \text{ kPa}, T_1 = 300 \text{ K}, P_2 = 150 \text{ kPa}, T_2 = 400 \text{ K}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$V_2 = \frac{P_1 V_1 T_2}{P_2 T_1} = \frac{(100 \text{ kPa})(100 \text{ mL})(400 \text{ K})}{(150 \text{ kPa})(300 \text{ K})} = 88.9 \text{ mL}$$

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CO Houle



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