

Unit 5 Types of Chemical Reactions Lab Report (37 pts)

In this experiment, you will mix chemicals and observe synthesis, single displacement, and double displacement reactions. You will make observations about what happens when you combine the different chemicals. Have your solubility chart. You will write balanced chemical equations for each reaction. You will need screen shots of your results at the end so be sure take snips of your results and paste them in the end right from the start.

This is in the content at Investigation: Types of Chemical Reactions 1 and you have to go to slide two to get to it:

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Investigation: Types of Chemical Reactions 1 ~

(3 points)

Part 1 Synthesis combustion reaction:

- You burned a Magnesium (Mg(s)) strip in the presence of Oxygen ($\text{O}_2\text{(g)}$) gas. Mg^{2+} , O^{2-}
 - What happened, describe what you saw, what happened in terms of left-over product, and add snip.
 - Write a **balanced** synthesis reaction for this process. I should see reactants < products
 - Using the A, B symbols, write a general equation for a synthesis reaction.

(4 points)

Part 2: Single-Displacement Reactions:

- For each of the four single-displacement reactions, describe what happened in each beaker. Use the Reactivity chart on page two to confirm your reaction.

Table 1 Observations Single Displacement Reactions			
Beaker	Solids (s)	Solutions (aq) 10 drops	Observations Before & After can add a screen shot
A	Aluminum	Copper sulfate $\text{CuSO}_4\text{(aq)}$	Darker specks develop on aluminum, this is copper
B	Zinc	Copper sulfate $\text{CuSO}_4\text{(aq)}$	
C	Zinc	Silver nitrate $\text{AgNO}_3\text{(aq)}$	
D	Copper	Silver nitrate $\text{AgNO}_3\text{(aq)}$	

- (1 pt) Using the A, B and C symbols, write a general equation for a single-displacement reaction.

(6 pts for products, balanced and parenthesis for solid, liquid, gas or aqueous)

- Write a **balanced** equation each reaction.

Ionic charges: Zn^{2+} , Al^{3+} , Cu^{2+} , Ag^{+} , SO_4^{2-} , NO_3^{-}

Here are the chemical formulas for the reactants in each reaction: I should see reactants < products. The reactants are given, but you need to cross down your oxidation numbers to get the correct chemical formulas for the products. Be sure to include (aq), or (s) for each chemical formula. I'll do the reaction A, you do the rest.

- A. Aluminum – Al(s) and copper sulfate – $\text{CuSO}_4\text{(aq)}$

1. If the charges of the ions are the same, then the ratio is 1:1.
2. If the charges of the ions are different, then the ratio can be determined by criss-crossing the charges.



Lab Report Types Of Chemical Reactions Answers

Ralph Thomas Myers



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