

Chapter 9 Practice Test - Naming and Writing Chemical Formulas

Matching

Match each item with the correct statement below.

Match each item with the correct statement below.

- | | |
|--------------------------------|--------------------|
| a. monatomic ion | f. cation |
| b. acid | g. binary compound |
| c. base | h. anion |
| d. law of definite proportions | i. polyatomic ion |
| e. law of multiple proportions | |

- _____ 1. consists of a single atom with a positive or negative charge
- _____ 2. atom or group of atoms having a negative charge
- _____ 3. atom or group of atoms having a positive charge
- _____ 4. tightly-bound group of atoms that behaves as a unit and carries a net charge
- _____ 5. compound composed of two different elements
- _____ 6. produces a hydrogen ion when dissolved in water
- _____ 7. produces a hydroxide ion when dissolved in water
- _____ 8. In any chemical compound, the masses of elements are always in the same proportion by mass.
- _____ 9. when two elements form more than one compound, the masses of one element that combine with the same mass of the other element are in the ratio of small, whole numbers

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 10. When naming a transition metal ion that can have more than one common ionic charge, the numerical value of the charge is indicated by a _____.
 - a. prefix
 - b. suffix
 - c. Roman numeral following the name
 - d. superscript after the name
- _____ 11. Which of the following correctly provides the name of the element, the symbol for the ion, and the name of the ion?
 - a. fluorine, F^+ , fluoride ion
 - b. zinc, Zn^{2+} , zincate ion
 - c. copper, Cu^+ , cuprous ion
 - d. sulfur, S^{2-} , sulfurous ion
- _____ 12. The nonmetals in Groups 6A and 7A _____.
 - a. lose electrons when they form ions
 - b. have a numerical charge that is found by subtracting 8 from the group number
 - c. all have ions with a -1 charge
 - d. end in *-ite*
- _____ 13. Which of the following is NOT a cation?
 - a. iron(III) ion
 - b. sulfate
 - c. Ca^{2+}
 - d. mercurous ion

Chapter 9 Practice Test Naming And Writing Chemical Formulas

Christian Drosten



Chapter 9 Practice Test Naming And Writing Chemical Formulas:

Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow, 1978 **Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson** Norman E.

Griswold, Henry Fuller Holtzclaw, 1988 *Holtzclaw Gen Chem Sg 9ed* Holtzclaw, 1991 Robinson Chemistry Study Guide

Robinson, 1992 **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Basic Chemical Principles** Edward I. Peters, 1988

College Study Skills Dianna L. Van Blerkom, 2000 This comprehensive text workbook focuses on getting students actively involved in the learning process and in learning how to learn Rather than presenting one prescriptive approach students take a learning style inventory and are then exposed to strategies that complement their personal learning styles

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