

## Chemical Equilibrium

**Ex 1.** For the equilibrium:  $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$

the equilibrium constant  $K_c$  is 50.5 at 448°C.

If  $[\text{H}_2]_0 = [\text{I}_2]_0 = 1 \text{ mol.L}^{-1}$ , what will the concentration of  $\text{H}_2$ ,  $\text{I}_2$  and  $\text{HI}$  be at equilibrium.

**Ex 2.** Consider the following reversible reaction:



One mole of  $\text{PCl}_5$  is placed in a flask of capacity 4 liter to be dissociated at constant temperature T. the number of moles of  $\text{Cl}_2$  formed at equilibrium is 0.25mole.

Calculate the equilibrium constant of the above system at temperature T.

**Ex 3.** At 100°C, the equilibrium constant for the reaction system,



is  $K_c = 2.19 \times 10^{-10}$

1. Determine whether the following mixture of components of the above system is at equilibrium:

$$[\text{COCl}_2] = 3.5 \times 10^{-3} \text{ mol.L}^{-1}$$

$$[\text{CO}] = 1.2 \times 10^{-5} \text{ mol.L}^{-1}$$

$$[\text{Cl}_2] = 3.3 \times 10^{-6} \text{ mol.L}^{-1}$$

2. If it is not at equilibrium, in what direction the system proceeds to reach equilibrium?

**Ex 4.** In a closed container of volume V and at constant temperature  $T = 300\text{K}$ , we introduce 2 mol of NO and 1 mol of  $\text{O}_2$  gases. the following equilibrium is established:



Calculate the degree of conversion ( $\alpha$ ) of NO if the amount of  $\text{NO}_2$  at equilibrium is three time that of  $\text{O}_2$ .

**Ex 5.** Ethanamine  $\text{C}_2\text{H}_5\text{NH}_2$  gas is put in an empty flask of volume  $V = 2\text{L}$ , maintained at constant temperature  $T = 500\text{K}$ .

The following equilibrium is established:



The initial amount of ethanamine is 0.2 mol and the total amount of the mixture at equilibrium is 0.3 mol.

1. Calculate the degree of dissociation  $\alpha$  of ethanamine.
2. Calculate the equilibrium constant  $K_c$  at 500K.

# Chemical Equilibrium Problems With Solutions

**Iris Wai-yin Shiu**



## **Chemical Equilibrium Problems With Solutions:**

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*Chemical Equilibrium* William Guenther, 2012-12-06 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical

methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations. The coverage of solution equilibrium phenomena is not however exhaustively comprehensive. Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria. Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation.

*Solving Equilibrium Problems with Applications to Qualitative Analysis* Steven S. Zumdahl, 1989

*Element Variables and the Solution of Complex Chemical Equilibrium Problems* Keith Meintjes, Alexander P. Morgan, 1987

**Principles and Applications of Aquatic Chemistry** François M. M. Morel, Janet G. Hering, 1993-03-10

Presents aquatic chemistry in a way that is truly useful to those with diverse backgrounds in the sciences. Major improvements to this edition include a complete rewrite of the first three background chapters making them user friendly. There is less emphasis on mathematics and concepts are illustrated with actual examples to facilitate understanding.

Equilibrium Calculations John T. Donoghue, 1971

*Tearing Versus Simultaneous Solution in Solving Chemical Equilibrium Problems* Iris Wai-yin Shiu, 1980

Exploring Chemical Analysis Daniel C. Harris, 2005

Exploring Chemical Analysis teaches students how to understand analytical results and how to use quantitative manipulations preparing them for the problems they will encounter.

**Surface Complexation Modeling** David A. Dzombak, François M. M. Morel, 1991-01-16

Provides a description of the thermodynamic model data treatment procedures and the thermodynamic constants for hydrous ferric oxide. Includes detailed coverage of the model and the parameter extraction procedure.

*A Numerical Algorithm for the Solution of Chemical Equilibrium Problems* Thomas Wayne Wall, 1984

Quantitative Chemical Analysis Mr. Rohit Manglik, 2024-07-11

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*Chemical Equilibrium* Allen J. Bard, 1966

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**Memorandum** Rand Corporation,1970      Study Questions and Problems in Inorganic Chemistry Alexander Silverman,1924      *Bulletin of the Operations Research Society of America* ,1972

**Thermodynamics and Chemistry, by F. H. MacDougall** Frank Henry Macdougall,1921      Journal of the American Chemical Society American Chemical Society,1914 Proceedings of the Society are included in v 1 59 1879 1937      General Catalogue Massachusetts Institute of Technology,1928      **Programme of the Courses of Instruction** Massachusetts Institute of Technology,1909

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