For a first order reaction,  $t_{1/2} = \frac{0.6932}{V}$  $k = \frac{0.6932}{6.932 \times 10^4}$  $k = 10^{-5} s^{-1}$  $\mathbf{k} = \begin{pmatrix} \frac{1}{t} \end{pmatrix} \ln \begin{pmatrix} \frac{[A_o]}{[A]} \end{pmatrix}$  $10^{-5} \text{ s}^{-1} \times 100 \times 60 \text{ s} = \ln \left( \frac{[A_0]}{[A]} \right)$  $0.06 = \ln \left( \frac{[A_0]}{[A]} \right)$  $\frac{[A_o]}{[A]} = e^{0.06}$  $\frac{[A_o]}{[A]} = 1.06$  $\therefore \frac{[A_o] - [A]}{[A]} \times 100 \%$  $= \left(1 - \frac{[A]}{[A]}\right) \times 100 \%$  $=\left(1-\frac{1}{1.06}\right)\times 100\%$ 

# **Chemistry Chemical Kinetics Problems Answers**

**Yicheng Fang** 

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