

Empirical and Molecular Formulas Worksheet

Key

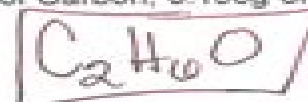
Objectives:

- be able to calculate empirical and molecular formulas

Empirical Formula

- 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

$$\begin{array}{l} \text{C} \quad \frac{0.783\text{g C}}{12.01\text{g}} = \frac{0.065}{0.033} = 1.96 \quad 2 \\ \text{H} \quad \frac{0.196\text{g H}}{1.01\text{g}} = \frac{0.194}{0.033} = 5.88 \quad 6 \\ \text{O} \quad \frac{0.521\text{g O}}{16.00\text{g}} = \frac{0.033}{0.033} = 1 \quad 1 \end{array}$$



- 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?

$$\text{Au} \quad \frac{89.14\text{g}}{196.97} = \frac{0.45}{0.45} = 1 \quad \text{AuO}_{1.5}$$

$$\text{O} \quad \frac{10.80\text{g}}{16\text{g}} = \frac{0.675}{0.45} = 1.5$$



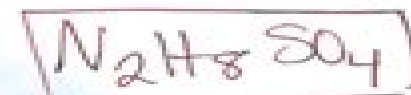
- 3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

$$\text{N} \quad \frac{21.2\text{g}}{14.01} = 1.5 \quad 2$$

$$\text{H} \quad \frac{6.1\text{g}}{1.01\text{g}} = 6.03 \quad 8$$

$$\text{S} \quad \frac{24.2\text{g}}{32} = 0.75 \quad 1$$

$$\text{O} \quad \frac{48.5\text{g}}{16.00\text{g}} = 3.03 \quad 4$$



Molecular Formula

- 4) Empirical formula of a substance is CH_2O . Molar mass is 180. What is the molecular formula?

$$\text{EFM} = (12 + 2 \times 1 + 16) =$$

$$\frac{180}{30} = 6$$



- 5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

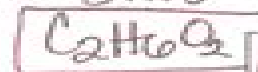
$$\text{C} \quad \frac{1.388}{3.585} \quad 38.7\% \rightarrow \frac{3.88\text{g}}{12.01\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{H} \quad \frac{0.345}{3.585} \quad \frac{9.62\text{g}}{1.01\text{g}} = \frac{9.52}{3.22} = 3$$

$$\text{O} \quad \frac{1.850}{3.585} \quad \frac{51.6\text{g}}{16\text{g}} = \frac{3.22}{3.22} = 1$$

$$\text{CH}_3\text{O} = \text{EFM} (12 + 3 \times 1 + 16)$$

$$\frac{62}{31.03} = 2$$



1	$\text{C}_2\text{H}_6\text{O}$
2	Au_2O_3
3	$\text{N}_2\text{H}_8\text{S}_4\text{O}_4$
4	$\text{C}_6\text{H}_{12}\text{O}_6$
5	$\text{C}_2\text{H}_6\text{O}_2$

Chemistry Empirical Formula Worksheet Answers

Richard Lloyd Whiteley



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Educart ICSE Class 10 One-shot Question Bank 2026 Chemistry (strictly for 2025-26 boards) Sir Tarun Rupani, 2025-07-12 Fast track your Chemistry revision with this exam ready resource This One shot Question Bank by Sir Tarun Rupani is designed to help ICSE Class 10 students revise the complete Chemistry syllabus quickly and thoroughly It simplifies theory boosts numerical accuracy and ensures strong exam practice all aligned with the 2025 26 ICSE syllabus Key Features Strictly Based on ICSE 2025 26 Curriculum Complete chapter coverage including Periodic Table Chemical Bonding Acid Base Organic Chemistry and more One shot Format Each chapter includes concise concept notes chemical equations reactions and key diagrams for quick recall Complete Coverage of Question Types Includes objective short long answers equation based numerical and reasoning questions Chapterwise PYQs Included Practice with previous years ICSE board questions to understand trends and improve retention Solved Answers in ICSE Format Clear well structured solutions using proper units chemical symbols and balanced equations Smart Revision Focus Special tips to avoid common mistakes in writing reactions balancing equations and attempting numericals Why Choose This Book This Chemistry One shot by Sir Tarun Rupani is built for smart preparation whether you re revising at the last minute or practising throughout the term It helps you approach each question with clarity confidence and the precision needed to score high in the 2026 ICSE board exam Objective Workbook for Simplified ICSE Chemistry , *Using ICT to Enhance Teaching and Learning in Chemistry* Steve Lewis, 2004 Information and Communications Technology has revolutionised the ways we process access and use information and as computers and other devices become ever more powerful and information becomes more readily available the next generation will need to be able to interact with digital media effectively to exploit these amazing new technologies to their full potential for the benefit of society It is therefore imperative that teachers become familiar with ICT and its true potential and can present information with a perspective similar to that which the present generation of young people is using to develop their interests in their everyday life This resource from the RSC gives teachers of chemistry the practical help needed to integrate ICT into their teaching and stimulate the enthusiasm of a new generation of scientists in the exciting new areas of chemistry that are opening up such as Nanoscience and Nanotechnology Furthermore it will be highly effective in developing the new ethos of sustainability that will be a major driving force behind the next advances in chemistry that are vital if we are to survive the manifold problems confronting society in the next few decades This resource is ideal for all secondary teachers of chemistry trainee teachers and university lecturers *Solving Empirical Formula Problems* , 2016 Continue your consideration of classic chemistry problems with a look at empirical formulas and examine how empirical formulas relate to molecular formulas *Chemistry, Problem-Solving Worktext* James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the

changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet **Quick**

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Long,Harold Victor Anderson,1928 Chemical calculations Richard Lloyd Whiteley,1892 *Chemistry* Neil D.

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Makes Perfect Chemistry Review and Workbook, Second Edition Marian DeWane,Heather Hattori,2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you ve learned along with a complete answer key to check your work

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