

# MANAGERIAL ACCOUNTING

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**Léo Pomar**



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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula  $m_{mix} = \frac{m_1 + m_2 + m_3}{V_1 + V_2 + V_3}$  1 5 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web determine the molarity for each of the following solutions 0 444 mol of  $\text{CoCl}_2$  in 0 654 l of solution 98 0 g of phosphoric acid  $\text{H}_3\text{PO}_4$  in 1 00 l of solution 0 2074 g of calcium hydroxide  $\text{Ca(OH)}_2$  in 40 00 ml of solution 10 5 kg

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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