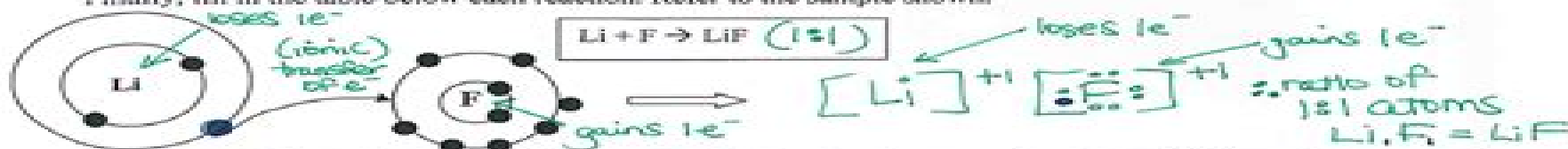


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1ve	transfers 1ve ⁻ to F	Li^+
F	7ve	accepts 1ve ⁻ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ 	Li	1ve	donate $1e^-$	Li^+
	Cl	1ve	accepts $1e^-$	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ 	Ca	2ve	donates $2e^-$	Ca^{+2}
	O	6ve	accepts $2e^-$	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ 	Be	2ve	donates $2e^-$ to each F	Be^{+2}
	F	7ve	each F accepts 1ve	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ 	Mg	2ve	donates $2e^-$ to S	Mg^{+2}
	S	6ve	accepts $2e^-$	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ 	K	1ve	donates 1ve ⁻	K^+
	F	7ve	accepts 1ve ⁻	F^-

Holt Chemistry Ionic Bonding Salts Answers

Harry G. Hajian, Robert L. Pecsok



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1. Understanding the eBook Holt Chemistry Ionic Bonding Salts Answers
 - The Rise of Digital Reading Holt Chemistry Ionic Bonding Salts Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Ionic Bonding Salts Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Ionic Bonding Salts Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Ionic Bonding Salts Answers
 - Personalized Recommendations
 - Holt Chemistry Ionic Bonding Salts Answers User Reviews and Ratings

- Holt Chemistry Ionic Bonding Salts Answers and Bestseller Lists
- 5. Accessing Holt Chemistry Ionic Bonding Salts Answers Free and Paid eBooks
 - Holt Chemistry Ionic Bonding Salts Answers Public Domain eBooks
 - Holt Chemistry Ionic Bonding Salts Answers eBook Subscription Services
 - Holt Chemistry Ionic Bonding Salts Answers Budget-Friendly Options
- 6. Navigating Holt Chemistry Ionic Bonding Salts Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Ionic Bonding Salts Answers Compatibility with Devices
 - Holt Chemistry Ionic Bonding Salts Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Ionic Bonding Salts Answers
 - Highlighting and Note-Taking Holt Chemistry Ionic Bonding Salts Answers
 - Interactive Elements Holt Chemistry Ionic Bonding Salts Answers
- 8. Staying Engaged with Holt Chemistry Ionic Bonding Salts Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Ionic Bonding Salts Answers
- 9. Balancing eBooks and Physical Books Holt Chemistry Ionic Bonding Salts Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Ionic Bonding Salts Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Holt Chemistry Ionic Bonding Salts Answers
 - Setting Reading Goals Holt Chemistry Ionic Bonding Salts Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Holt Chemistry Ionic Bonding Salts Answers
 - Fact-Checking eBook Content of Holt Chemistry Ionic Bonding Salts Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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