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Factor	Effect on rate	Diagram	Rate of reaction	Rate of reaction
Temperature	Increases		Rate of reaction increases	Rate of reaction increases
Concentration	Increases		Rate of reaction increases	Rate of reaction increases

10. Explain the effect of the following factors on the rate of reaction.

(a) Temperature: The rate of reaction increases with temperature. This is because the molecules have more kinetic energy and move faster, leading to more frequent collisions. The rate of reaction also increases with the surface area of the reactants.

(b) Concentration: The rate of reaction increases with the concentration of the reactants. This is because there are more molecules in a given volume, leading to more frequent collisions. The rate of reaction also increases with the surface area of the reactants.

(c) Surface Area: The rate of reaction increases with the surface area of the reactants. This is because there are more molecules exposed to the other reactant, leading to more frequent collisions. The rate of reaction also increases with the concentration of the reactants.

(d) Catalyst: A catalyst is a substance that speeds up the rate of reaction without being consumed in the process. It does this by providing an alternative pathway for the reaction that has a lower activation energy. The rate of reaction also increases with the surface area of the reactants.

(e) Pressure: The rate of reaction increases with the pressure of the reactants. This is because there are more molecules in a given volume, leading to more frequent collisions. The rate of reaction also increases with the surface area of the reactants.

(f) Light: The rate of reaction increases with the intensity of light. This is because the light provides energy to the molecules, leading to more frequent collisions. The rate of reaction also increases with the surface area of the reactants.

(g) pH: The rate of reaction increases with the pH of the reactants. This is because the pH affects the concentration of the reactants, leading to more frequent collisions. The rate of reaction also increases with the surface area of the reactants.

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