

Simple Thermodynamic Systems

2.1 THERMODYNAMIC EQUILIBRIUM

Suppose that experiments have been performed on a thermodynamic system and that the coordinates necessary and sufficient for a macroscopic description have been determined. When these coordinates change in any way whatsoever, either spontaneously or by virtue of outside influence, the system is said to undergo a *change of state*.[†] When a system is not influenced in any way by its surroundings, it is said to be *isolated*. In practical applications of thermodynamics, isolated systems are of little importance. We usually have to deal with a system that is influenced in some way by its surroundings. In general, the surroundings may exert forces on the system or provide contact between the system and a body at some definite temperature. When the state of a system changes, interactions usually take place between the system and its surroundings.

When there is no unbalanced force or torque in the interior of a system and also none between a system and its surroundings, the system is said to be in a state of *mechanical equilibrium*. When these conditions are not satisfied, either the system alone or both the system and its surroundings will undergo a change of state, which will cease only when mechanical equilibrium is restored.

When a system in mechanical equilibrium does *not* tend to undergo a spontaneous change of internal structure, such as a chemical reaction, or a transfer of matter from one part of the system to another, such as diffusion or solution, however slow, then it is said to be in a state of *chemical equilibrium*.

[†] Notice that in general physics the expression "change of state" is often used to signify a transition from solid to liquid or liquid to gas, etc. Such a change in the language of thermodynamics is called a *change of phase*.

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