Carbonate Alkalinity

 Consider a solution containing a fixed total dissolved concentration of CaCO₃. At the CO₂ equivalence point, H₂CO₃ is the principal carbonate species, so we choose our components as H⁺, H₂O, CO₂, and Ca²⁺.

$$H_2CO_3 = H_2O + CO_2$$

 $HCO_3^- = H_2O + CO_2 - H^+$
 $CO_3^{2-} = H_2O + CO_2 - 2H^+$

The proton mole balance equation is then:

$$TOTH = [H^+] - [HCO_3^-] - 2[CO_3^{2-}] - [OH^-]$$

The alkalinity is then:

$$Alk = -TOTH = -[H^+] + [HCO_3^-] + 2[CO_3^{2-}] + [OH^-]$$

- The above is also the carbonate alkalinity: the alkalinity due to the presence of carbonate ions.
 - In many solutions, carbonate alkalinity is very nearly equal total alkalinity. To be clear, we sometimes call the alkalinity as defined previously Total Alkalinity.

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Lingjun Ying

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