CHM 140 Lab 5 Synthesis of Alum from Aluminum Foil

Purpose: In this experiment, a commercially useful compound, alum, is synthesized from aluminum foil, KOH, and H,SO, according to the following balanced reaction:

Calculating the moles of each reactant used in the synthesis will allow for the determination of the limiting reactant. Once the limiting reactant is determined, the theoretical yield of 2KAl(SO₂)₂(H₂O)₁₂ can be determined. Then, the percent yield can be calculated by dividing the actual yield by the theoretical yield and multiplying by 100.

Data from the video will be used for the calculations.

https://www.be/K9F2ZUpnshl Synthesis of Alum

- 1. Mass of aluminum
- 2. Mass of empty Buchner funnel and filter paper
- 3. Mass of Buchner funnel, filter paper, and alum
- 4. Mass of alum (actual yield)
- 5. Volume of 1.4 M KOH used
- 6. Volume of 9.0 M H,SO, used

0.545	8
16.153	. 8
71977	-

8 919 e

12.0 ml

Calculations: Determine the limiting reactant. Show work for calculations.

7. Moles of Al

moles Al

8. Moles of KOH

9. Moles of H,SO,

moles KOH

moles H_iSO_i

Alum From Aluminum Lab Answers

Xiang Xie

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Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

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Technical Association of the Pulp and Paper Industry ,1955 History, Geology, and Environmental Setting of the Lower Fair Day Mine, Arapaho/Roosevelt National Forest, Boulder County, Colorado John Neubert,2000 Nebraska Experiment Station Quarterly ,1952 Plating and Surface Finishing ,2000 Chemical, Color and Oil Record ,1923

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Table of Contents Alum From Aluminum Lab Answers

- 1. Understanding the eBook Alum From Aluminum Lab Answers
 - The Rise of Digital Reading Alum From Aluminum Lab Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Alum From Aluminum Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Alum From Aluminum Lab Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Alum From Aluminum Lab Answers
 - Personalized Recommendations
 - Alum From Aluminum Lab Answers User Reviews and Ratings
 - Alum From Aluminum Lab Answers and Bestseller Lists
- 5. Accessing Alum From Aluminum Lab Answers Free and Paid eBooks
 - Alum From Aluminum Lab Answers Public Domain eBooks
 - Alum From Aluminum Lab Answers eBook Subscription Services
 - Alum From Aluminum Lab Answers Budget-Friendly Options
- 6. Navigating Alum From Aluminum Lab Answers eBook Formats

- o ePub, PDF, MOBI, and More
- Alum From Aluminum Lab Answers Compatibility with Devices
- Alum From Aluminum Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Alum From Aluminum Lab Answers
 - Highlighting and Note-Taking Alum From Aluminum Lab Answers
 - Interactive Elements Alum From Aluminum Lab Answers
- 8. Staying Engaged with Alum From Aluminum Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Alum From Aluminum Lab Answers
- 9. Balancing eBooks and Physical Books Alum From Aluminum Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Alum From Aluminum Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Alum From Aluminum Lab Answers
 - Setting Reading Goals Alum From Aluminum Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Alum From Aluminum Lab Answers
 - Fact-Checking eBook Content of Alum From Aluminum Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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