

Practice Problems: Quantum Theory

CHEM 1A

1. Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant = 6.626×10^{-34} J·s, speed of light = 2.998×10^8 m/s

$$E_{\text{photon}} = h\nu = \frac{hc}{\lambda}$$

$$E_{\text{photon}} = \frac{(6.626 \times 10^{-34} \text{ J}\cdot\text{s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$E_{\text{photon}} = \underline{3.02677861 \times 10^{-19} \text{ J}}$$

$$\nu = \frac{c}{\lambda}$$

$$\nu = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}} \right)}$$

$$\nu = \underline{4.568032912 \times 10^{14} \text{ s}^{-1}}$$

Answer (E): $3.027 \times 10^{-19} \text{ J}$ Answer (ν): $4.568 \times 10^{14} \text{ s}^{-1}$ or Hz

2. What is the de Broglie wavelength for an electron moving at 2.42×10^6 m/s?
Note: $m_e = 9.109 \times 10^{-31}$ kg

$$\lambda = \left(\frac{h}{mv} \right)$$

$$\lambda = \left(\frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{(9.109 \times 10^{-31} \text{ kg}) (2.42 \times 10^6 \text{ m/s})} \right) = \underline{3.00583657 \times 10^{-10} \text{ m}}$$

Answer: $3.01 \times 10^{-10} \text{ m}$ $\longleftarrow$ many orders of magnitude larger than the diameter of an electron

3. Using the Bohr equation, calculate is the change in energy when an electron transitions from $n = 3$ to $n = 2$ in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{initial}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} \text{ J} \left(\frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = \underline{-3.02777778 \times 10^{-19} \text{ J}}$$

Answer: $-3.03 \times 10^{-19} \text{ J}$

Quantum Theory And The Atom Worksheet Answers

Frank J. Bockhoff



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Introduction To The Quantum World Of Atoms And Molecules Jens Peder Dahl, 2001-06-15 This invaluable book provides a balanced and integrated introduction to the quantum world of atoms and molecules The underlying basis of quantum mechanics is carefully developed with respect for the historical tradition and from a molecular angle The fundamental concepts in the theory of atomic and molecular structure are thoroughly discussed as are the central techniques needed in quantum chemical applications Special attention is paid to exposing and clarifying the common ground of Hartree Fock theory and density functional theory Throughout the text the discussion is pedagogically obliging and aims at simplicity and mathematical clarity while avoiding the use of advanced mathematics End of chapter problems supplement the main text

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examples is also provided The emphasis throughout is on concepts References to experiments are made to test and support theories and introduce new information Mathematical complexity has been avoided as much as possible and problems of both a formal analytical and numerical kind are Quantum Mechanics of Atomic Spectra and Atomic Structure Masataka Mizushima,1970 Problems after each chapter **The Quantum Theory of the Atom** George Birt histle,1926

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