

6. When one gram acid of KMnO_4 reacts with HCl , the volume of chlorine liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres
(c) 44.8 litres (d) 56.8 litres
7. When MnO_4^- ions reacts with Fe^{2+} ions in acid solution the Fe^{3+} is liberated. The equivalent weight of KMnO_4 in this reaction is
- (a) 500 (b) 1500
(c) 300 (d) 2000
8. 54 g. of hydrogen peroxide is present in 1000 ml. of solution. This solution is called
- (a) 10 vol solution (b) 20 vol solution
(c) 30 vol solution (d) 50 vol solution
9. The number of moles of KMnO_4 that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 10 (b) 20
(c) 40 (d) 2
10. To prepare a solution that is 0.05 M KCl starting with 100 mL of 0.40 M KCl
- (a) add 0.70 g KCl
(b) add 50 mL of water
(c) add 0.70 mol KCl
(d) evaporate 10 mL water
11. In hot alkaline solution, Br_2 disproportionates to Br^- and BrO_2^-
- $$3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_2^- + 3\text{H}_2\text{O}$$
- Atomic equivalent weight of Br_2 is (molecular weight = 160)
- (a) $\frac{16}{5}$ (b) $\frac{16}{3}$
(c) $\frac{160}{5}$ (d) $\frac{160}{3}$
12. (Fe^{2+}) in a solution prepared by adding 20.00 mL of 0.10 M FeSO_4 with 70 mL of 0.10 M Na_2SO_4 is
- (a) 0.020 M (b) 0.140 M
(c) 0.010 M (d) 0.200 M
13. When 50 mL of 0.100 M HCl is mixed with 100 mL of 0.10 M KOH , the resulting solution is the same as a solution of
- (a) 0.10 M HCl and 0.05 M HCl
(b) 0.05 M HCl
(c) 0.05 M HCl and 0.05 M KOH
(d) 0.05 M HCl and 0.05 M HCl
14. 50 mL of H_2O_2 solution contains strength = x% required 10 mL of KMnO_4 solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.4
(c) 0.5 (d) 06.8
15. Molarity of 1.0 M H_2SO_4 (d = 1.8 mL/cm³) is
- (a) 36 mol kg⁻¹ (b) 300 mol kg⁻¹

Questions With Answers Inorganic Chemistry Multiple

**Prof. S. K. Khanna, Dr. N. K. Verma, Dr.
B. Kapila**



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Inorganic Chemistry Ram Charitra Maurya, 2021-04-06 This book covers different aspects of Inorganic Chemistry in terms of 10 Chapters with in depth and up to date coverage Starting with the VSEPR theory in the first chapter the book symmetrically presents delocalized p bonding in polyatomic molecules structure bonding and topology of borane and related compounds synthesis and reactivity of metal clusters and their bonding some aspects of stability constants of metal complexes magnetochemistry mechanism of inorganic reactions molecular orbital MO approach of bonding in transition metals bonding in organometallic sandwich compounds based on MO approach Safe and economical inorganic experiments at UG and PG Levels are also presented in the last chapter At the end five relevant topics are included as appendices for updating students and faculty members

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ExamCrackers MCAT. Jonathan Orsay,2005 *CBSE Class 12 Chemistry Handbook - MINDMAPS, Solved Papers, Objective Question Bank & Practice Papers* Disha Experts,2019-07-19

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Frontier Computing Jason C. Hung,Neil Y. Yen,Jia-Wei Chang,2023-05-24 This book

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