

Simple Thermodynamic Systems

2.1 THERMODYNAMIC EQUILIBRIUM

Suppose that experiments have been performed on a thermodynamic system and that the coordinates necessary and sufficient for a macroscopic description have been determined. When these coordinates change in any way whatsoever, either spontaneously or by virtue of outside influence, the system is said to undergo a *change of state*.[†] When a system is not influenced in any way by its surroundings, it is said to be *isolated*. In practical applications of thermodynamics, isolated systems are of little importance. We usually have to deal with a system that is influenced in some way by its surroundings. In general, the surroundings may exert forces on the system or provide contact between the system and a body at some definite temperature. When the state of a system changes, interactions usually take place between the system and its surroundings.

When there is no unbalanced force or torque in the interior of a system and also none between a system and its surroundings, the system is said to be in a state of *mechanical equilibrium*. When these conditions are not satisfied, either the system alone or both the system and its surroundings will undergo a change of state, which will cease only when mechanical equilibrium is restored.

When a system in mechanical equilibrium does *not* tend to undergo a spontaneous change of internal structure, such as a chemical reaction, or a transfer of matter from one part of the system to another, such as diffusion or solution, however slow, then it is said to be in a state of *chemical equilibrium*.

[†] Notice that in general physics the expression "change of state" is often used to signify a transition from solid to liquid or liquid to gas, etc. Such a change in the language of thermodynamics is called a *change of phase*.

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 - The Rise of Digital Reading Heat Y Thermodynamics Zemansky Solutions Bing
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Popular eBook Platforms
 - Features to Look for in an Heat Y Thermodynamics Zemansky Solutions Bing
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heat Y Thermodynamics Zemansky Solutions Bing
 - Personalized Recommendations
 - Heat Y Thermodynamics Zemansky Solutions Bing User Reviews and Ratings

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 - ePub, PDF, MOBI, and More
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 - Adjustable Fonts and Text Sizes of Heat Y Thermodynamics Zemansky Solutions Bing
 - Highlighting and Note-Taking Heat Y Thermodynamics Zemansky Solutions Bing
 - Interactive Elements Heat Y Thermodynamics Zemansky Solutions Bing
- 8. Staying Engaged with Heat Y Thermodynamics Zemansky Solutions Bing
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Heat Y Thermodynamics Zemansky Solutions Bing
- 9. Balancing eBooks and Physical Books Heat Y Thermodynamics Zemansky Solutions Bing
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heat Y Thermodynamics Zemansky Solutions Bing
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Heat Y Thermodynamics Zemansky Solutions Bing
 - Setting Reading Goals Heat Y Thermodynamics Zemansky Solutions Bing
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Heat Y Thermodynamics Zemansky Solutions Bing
 - Distinguishing Credible Sources

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 - Exploring Educational eBooks
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 - Interactive and Gamified eBooks

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