

SECTION : 3 (Maximum Marks : 16)

- This section contains TWO paragraphs
- Based on each paragraph, there will be TWO questions.
- Each question has FOUR options (A), (B), (C) and (D). ONE OR MORE THAN ONE of these four option(s) is(are) correct.
- For each question, darken the bubble(s) corresponding to all the correct option(s) in the ORS.
- Marking scheme:
 +4 If only the bubble(s) corresponding to all the correct option(s) is(arc) darkened.
 0 If one of the bubbles is darkened.
 -2 In all other cases.

PARAGRAPH – 1

When 100 mL of 1.0 M HCl was mixed with 100 mL of 1.0 M NaOH in an insulated beaker at constant pressure, a temperature increase of 5.7°C was measured for the beaker and its contents (Expt. 1). Because the enthalpy of neutralization of a strong acid with a strong base is a constant ($-57.0 \text{ kJ mol}^{-1}$), this experiment could be used to measure the calorimeter constant. In a second experiment (Experiment 2) 100 mL of 2.0 M acetic acid ($K_a = 2.0 \times 10^{-5}$) was mixed with 100 mL of 1.0 M NaOH (under identical conditions of Expt. 1) where a temperature rise of 5.6°C was measured.

(Consider heat capacity of all solutions as $4.2 \text{ J g}^{-1} \text{ K}^{-1}$ and density of solutions as 1.0 g mL^{-1})

37. Enthalpy of dissociation (in kJ mol^{-1}) of acetic acid obtained from the Expt. 2 is
 (A) 1.0 (B) 10.0 (C) 24.5 (D) 51.4

KEY: (A)

SOL: Standard enthalpy of neutralization = -57 kJ/mol

Amount of acid used = 0.1 mole

Amount of base used = 0.1 mole

Amount of heat released on neutralization = 5.7 kJ

Change in temperature = 5.7°C

Let the calorimeter constant be x

Now, $5700 \text{ J} = 200 \times 4.2 \times 5.7 + x \times 5.7$

$$x = 160 \text{ J}$$

For acetic acid and NaOH.

Mole of acetic acid neutralized = 0.1

Energy released in neutralization = $200 \times 5.6 \times 4.2 + 160 \times 5.6 = 5600 \text{ J}$

Energy released per mole = $\frac{5600}{0.1} = 56000 \text{ J} = 56 \text{ kJ}$

Enthalpy of ionization of acetic acid = 1 kJ .

Hence answer is (A)

Jee Advanced 2015 Hints Solutions Paper 2 Narayana

Er. Sunil Batra



Jee Advanced 2015 Hints Solutions Paper 2 Narayana:

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