

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

losing

true

true

polyatomic

seven

true (released)

9. Calcium becomes a monatomic cation by gaining two electrons.

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

12. Monatomic ions consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

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Donald J. Guerrieri, Haber, Hoyt, Turner



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Table of Contents 16 1 Review And Reinforcement Answers Key

1. Understanding the eBook 16 1 Review And Reinforcement Answers Key
 - The Rise of Digital Reading 16 1 Review And Reinforcement Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 16 1 Review And Reinforcement Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 16 1 Review And Reinforcement Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 16 1 Review And Reinforcement Answers Key
 - Personalized Recommendations
 - 16 1 Review And Reinforcement Answers Key User Reviews and Ratings

- 16 1 Review And Reinforcement Answers Key and Bestseller Lists
- 5. Accessing 16 1 Review And Reinforcement Answers Key Free and Paid eBooks
 - 16 1 Review And Reinforcement Answers Key Public Domain eBooks
 - 16 1 Review And Reinforcement Answers Key eBook Subscription Services
 - 16 1 Review And Reinforcement Answers Key Budget-Friendly Options
- 6. Navigating 16 1 Review And Reinforcement Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 16 1 Review And Reinforcement Answers Key Compatibility with Devices
 - 16 1 Review And Reinforcement Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 16 1 Review And Reinforcement Answers Key
 - Highlighting and Note-Taking 16 1 Review And Reinforcement Answers Key
 - Interactive Elements 16 1 Review And Reinforcement Answers Key
- 8. Staying Engaged with 16 1 Review And Reinforcement Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 16 1 Review And Reinforcement Answers Key
- 9. Balancing eBooks and Physical Books 16 1 Review And Reinforcement Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 16 1 Review And Reinforcement Answers Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 16 1 Review And Reinforcement Answers Key
 - Setting Reading Goals 16 1 Review And Reinforcement Answers Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 16 1 Review And Reinforcement Answers Key
 - Fact-Checking eBook Content of 16 1 Review And Reinforcement Answers Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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