

Stoichiometry of a Precipitation Reaction

Hands-On Labs, Inc.

Version 42-0201-00-02

Lab Report Assistant

This document is not meant to be a substitute for a formal laboratory report. The Lab Report Assistant is simply a summary of the experiment's questions, diagrams if needed, and data tables that should be addressed in a formal lab report. The intent is to facilitate students' writing of lab reports by providing this information in an editable file which can be sent to an instructor.

Exercise 1: Stoichiometry and a Precipitation Reaction

Data Table 1. Stoichiometry Values.

Initial: $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ (g)	1.00
Initial: $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ (moles)	0.0068
Initial: CaCl_2 (moles)	0.008
Initial: Na_2CO_3 (moles)	0.0068
Initial: Na_2CO_3 (g)	0.7
Theoretical: CaCO_3 (g)	0.68g
Mass of Filter paper (g)	90g
Mass of Filter Paper + CaCO_3 (g)	1.00
Actual: CaCO_3 (g)	0.45
% Yield:	67.5 %

Questions

- A perfect percent yield would be 100%. Based on your results, describe your degree of accuracy, and suggest possible sources of error. My percent yield was 67.5 % and even though a perfect percent yield is 100 I was still in a good range. Possible sources of error could have been that my measurements were not as accurate as they needed to be and possibly that my calculations were off.
- What impact would be adding twice as much Na_2CO_3 than required for stoichiometric quantities have on the quantity of product produced? It could interfere with other reactions that are supposed to react with the Na_2CO_3 .
- Determine the quantity (g) of pure CaCl_2 in 7.5 g of $\text{CaCl}_2 \cdot 9\text{H}_2\text{O}$. 3.0

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