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Factor	Effect on rate	Diagram	Rate of reaction	Rate of reaction
Temperature	Increases		Rate of reaction increases	Rate of reaction increases
Concentration	Increases		Rate of reaction increases	Rate of reaction increases

14. Explain the effect of the following factors on the rate of reaction.

(a) Temperature: As temperature increases, the rate of reaction increases. This is because the molecules have more kinetic energy and are moving faster, so they are more likely to collide and react. The diagram shows a graph of rate of reaction versus temperature. The curve starts at the origin and rises steeply, then levels off.

(b) Concentration: As concentration increases, the rate of reaction increases. This is because there are more molecules in the same volume, so they are more likely to collide and react. The diagram shows a graph of rate of reaction versus concentration. The curve starts at the origin and rises steeply, then levels off.

(c) Surface area: As surface area increases, the rate of reaction increases. This is because there are more molecules exposed to the other reactant, so they are more likely to collide and react. The diagram shows a graph of rate of reaction versus surface area. The curve starts at the origin and rises steeply, then levels off.

(d) Catalyst: A catalyst increases the rate of reaction by providing an alternative pathway for the reaction. The diagram shows a graph of rate of reaction versus time. The curve with the catalyst starts at the origin and rises more steeply than the curve without the catalyst, reaching the same level of reaction faster.

(e) Pressure: As pressure increases, the rate of reaction increases. This is because the molecules are packed more closely together, so they are more likely to collide and react. The diagram shows a graph of rate of reaction versus pressure. The curve starts at the origin and rises steeply, then levels off.

(f) Light: Light can increase the rate of reaction in some cases, such as in the reaction between hydrogen peroxide and iodide ions. The diagram shows a graph of rate of reaction versus time. The curve with light starts at the origin and rises more steeply than the curve without light, reaching the same level of reaction faster.

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