

## 8.2 – THE CHEMICAL EARTH

### 1 – The living and non-living components of the Earth contain mixtures

#### 1.1 Construct word and balanced formulae equations of chemical reactions as they are encountered

- To make **word equations**, write each chemical formula, in words. Then, add in pluses and arrows.



- To create **balanced formulae equations**, write down **chemical symbols** (see 1.7), their **state of matter** and **balance** the equation.

|                                                                    |                                                                                                     |
|--------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------|
| Magnesium + Oxygen $\rightarrow$ Magnesium Oxide                   | The original word equation.                                                                         |
| $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$                    | There are not enough Oxygen atoms on the right.                                                     |
| $\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$                   | Double the <b>whole molecule</b> , not just magnesium.<br>There are not enough magnesium atoms now. |
| $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$                  | Add a 2 in front of magnesium on the left. Now, LHS = RHS                                           |
| $2\text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2\text{MgO}_{(s)}$ | Place the states of matter for each substance.                                                      |

- Combustion – **Fuel + Oxygen  $\rightarrow$  Carbon dioxide + Water**, example:  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- Corrosion – **Metal + Oxygen  $\rightarrow$  Metal Oxide**, example:  $2\text{Cu} + \text{O}_2 \rightarrow 2\text{CuO}$
- Respiration – **Glucose + Oxygen  $\rightarrow$  Carbon Dioxide + Water**, example:  $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$
- Precipitation – **Soluble salt + Soluble salt  $\rightarrow$  Insoluble salt + Soluble salt**
  - o Example:  $\text{CuSO}_4 + 2\text{NaOH} \rightarrow \text{Cu}(\text{OH})_2 + \text{Na}_2\text{SO}_4$  ( $\text{Cu}(\text{OH})_2$  is insoluble)
- Neutralisation – **Acid + Base  $\rightarrow$  Salt + Water**, example:  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- Acids on metals – **Metal + Acid  $\rightarrow$  Salt + Hydrogen**, example:  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$
- Acid on metal oxide – **Metal Oxide + Acid  $\rightarrow$  Salt + Water**, example:  $\text{MgO} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$
- Acids on carbonates – **Carbonate + Acid  $\rightarrow$  Salt + Water + Carbon Dioxide**
  - o Example:  $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$
- Decomposition – **Molecule  $\rightarrow$  Molecule part + Molecule part**, example:  $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$

#### 1.2 Identify the difference between elements, compounds and mixtures in terms of particle theory

- Particle theory helps us **visualise** differences between elements, compounds and mixtures.
- **Elements** are basic building blocks, having own unique symbol and physical and chemical properties.
- **Compounds** are **two or more particles** chemically bonded.
- A **mixture** is composed of a mix of **elements and compounds**, e.g. a solution.
  - o Can be separated using simple separation techniques, does not have fixed properties.

# Secondary 2 Chemistry Notes 2p115science Wikispaces

**Sebastian Brünink**



**Secondary 2 Chemistry Notes 2p115science Wikispaces :**

The book delves into Secondary 2 Chemistry Notes 2p115science Wikispaces . Secondary 2 Chemistry Notes 2p115science Wikispaces is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Secondary 2 Chemistry Notes 2p115science Wikispaces , encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Secondary 2 Chemistry Notes 2p115science Wikispaces
- Chapter 2: Essential Elements of Secondary 2 Chemistry Notes 2p115science Wikispaces
- Chapter 3: Secondary 2 Chemistry Notes 2p115science Wikispaces in Everyday Life
- Chapter 4: Secondary 2 Chemistry Notes 2p115science Wikispaces in Specific Contexts
- Chapter 5: Conclusion

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