

Colligative Properties and Activity Coefficients Exercise 1 and 2 (continued)

Knowledge being tested: to solve what depends on _____

Solution

For activity coefficients: $\gamma_{\pm}$

Using the following information:

Mean ionic activity coefficient for NaCl at 25°C and 0.1 mol kg^{-1} is $\gamma_{\pm} = 0.779$.

$$\gamma_{\pm} = \frac{\gamma_{\text{Na}^+} \gamma_{\text{Cl}^-}}{\gamma_{\text{NaCl}}} = \frac{\gamma_{\text{Na}^+} \gamma_{\text{Cl}^-}}{\gamma_{\text{NaCl}}}$$

Therefore:

$$\gamma_{\text{Na}^+} \gamma_{\text{Cl}^-} = \gamma_{\pm}^2 = (0.779)^2 = 0.607$$

Experimental Determination of Activity Coefficients

The paper presents the following information:

When a solution of a salt is prepared, the activity of the resulting solution is _____ (Figure 10.1, Table 10.1, the value of the activity of the pure solvent).

Colligative Properties Study Guide Answers

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